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# Influence of the microstructure and laser shock processing (LSP) on the corrosion behaviour of the AA2050-T8 aluminium alloy

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## A B S T R A C T

The corrosion behaviour of AA2050-T8 was studied after polishing and after laser shock processing (LSP) treatment using the electrochemical microcell technique and the SVET. After polishing, pitting at constituent particles and intergranular corrosion were observed. By contrast, no intergranular corrosion developed after LSP. Microcell measurements revealed that LSP increases the pitting potential. SVET measurements revealed that local anodic currents are systematically lower on LSP-treated surfaces than on polished ones. The current density on the LSP-treated surface remains constant around  $50 \mu\text{A cm}^{-2}$  up to 123 min after immersion, while on the polished surface it reaches  $200 \mu\text{A cm}^{-2}$ .

### Keywords:

A. Aluminium alloy

B. Polarization

C. Intergranular corrosion

C. Pitting corrosion

C. Effects of strain

## 1. Introduction

Aluminium alloys are a very important category of materials because of their outstanding mechanical properties and wide range of industrial applications. It is well known [1–3] that aluminium alloys are prone to microstructural corrosion (pitting corrosion, intergranular corrosion, etc.). Microstructural corrosion is caused by the presence of intermetallic particles, which may be either anodic or cathodic with respect to the matrix [1–4]. Micro-electrochemical techniques have been widely employed to obtain information about the role of these particles in localised corrosion of aluminium alloys. Park et al. [5] used a pH microelectrode and demonstrated that alkalisation occurred around a synthetic  $\text{Al}_3\text{Fe}$  ( $200 \times 200 \mu\text{m}^2$ ) electrode coupled to a 6061 alloy sample in NaCl 0.6 M at pH 6.3. The local dissolution of various aluminium alloys has been probed *in situ* [6] in chloride solutions by using atomic force microscopy (AFM) and scanning electrochemical microscopy (SECM). Preferential dissolution in the boundary region between some intermetallic particles (IMPs) and the alloy matrix, and trench formation around large IMPs during free immersion and under electrochemical anodic polarisation were observed. The use of microcapillary cell techniques [7–10] enables the measurement of local breakdown potentials for areas in aluminium alloys containing

one or a few intermetallics [11–13]. A mathematical model for simulating the local electrochemical response of a passive aluminium surface with an active pit has been developed [14]. This model was applied to a real capillary geometry used in electrochemical microcell experiments. It was found that for  $r_{\text{cap}}/r_{\text{pit}} < 100$ , where  $r_{\text{cap}}$  and  $r_{\text{pit}}$  are the capillary tip and pit radii respectively, the insulating capillary wall affects the species concentrations and the electrolyte potential. Moreover, for  $r_{\text{cap}}/r_{\text{pit}} < 20$ , the shape of the capillary, which might not be cylindrical, should be taken into account.

Main applications of LSP treatment are directed to aerospace and automotive industries [15]. LSP treatment consists in irradiating material surfaces with nanosecond laser pulses that generate plasma-driven shock waves, which in turn lead to a certain amount of local plastic deformation. Thermal effects are avoided by covering the surface with an absorptive overlay. Deep (i.e. greater than 0.5 mm) and high-amplitude compressive residual stress fields, combined with relatively low work hardening rates (less than 20%) and limited roughening, are usually the main characteristics of surfaces having undergone LSP treatment.

The beneficial effects of LSP on static, cyclic and fretting fatigue performance of aluminium alloys, steels and nickel-based alloys have been demonstrated [15 and Ref. therein]. LSP treatment is also expected to improve the corrosion behaviour of aluminium alloys. Previous works at the macro-scale have provided evidence of an anodic shift of the pitting potential of 2024-T351 aluminium alloy treated by LSP [16]. In this paper, the influence of the specimen microstructure and LSP on the resistance to localised corrosion of

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AA2050-T8 aluminium alloy is investigated in chloride media at the microscale using microcapillaries, the SVET and FE-SEM/EDS.

## 2. Experimental

### 2.1. Specimens and surface preparation

Experiments were performed on the 2050-T8 aluminium alloy (Cu: 3.73 wt.%; Li: 0.9 wt.%; Mg: 0.3 wt.%; Mn: 0.37 wt.%; Zr: 0.009 wt.% and Fe: 0.05 wt.%). Specimens were ground with silicon carbide (SiC) papers down to 4000 grit and polished with diamond pastes (9, 3 and 1  $\mu\text{m}$ ).

Prior to LSP, specimens were covered with an 80  $\mu\text{m}$ -thick metal adhesive coating so as to prevent any thermal rise during LSP and hence any microstructural changes. As shown in Fig. 1(a), the specimen surface was partially treated by LSP. Experiments were carried out with a frequency-doubled (0.532  $\mu\text{m}$ ) Nd:YAG laser capable of delivering up to 1.5 J in pulses of 9 ns duration at a 10 Hz repetition rate. 1.8 mm impact diameters were used to treat 1.5 cm circular areas. The intensity value was 3.5  $\text{GW cm}^{-2}$ .

### 2.2. Global and local electrochemical measurements

Global polarisation curves were determined on the polished and LSP-treated surfaces using a classical three-electrode cell. All potentials were measured vs. Ag/AgCl (in saturated KCl solution) and the counter electrode was a platinum disk (1 cm diameter). An AutoLab PGZ128 electrochemical interface was used in these experiments. On the other hand, local polarisation curves were determined using the electrochemical microcell technique [17–19]. The microcell was mounted on a microscope for precise positioning of the capillary on its surface. The capillary tip was sealed to the specimen surface with a layer of silicon rubber. The wetted area on the specimen could be calculated accurately using a specific image analysis procedure. All potentials were measured vs. Ag/AgCl (saturated KCl solution) and the counter electrode was a platinum wire. Local polarisation curves were determined using capillary diameters of

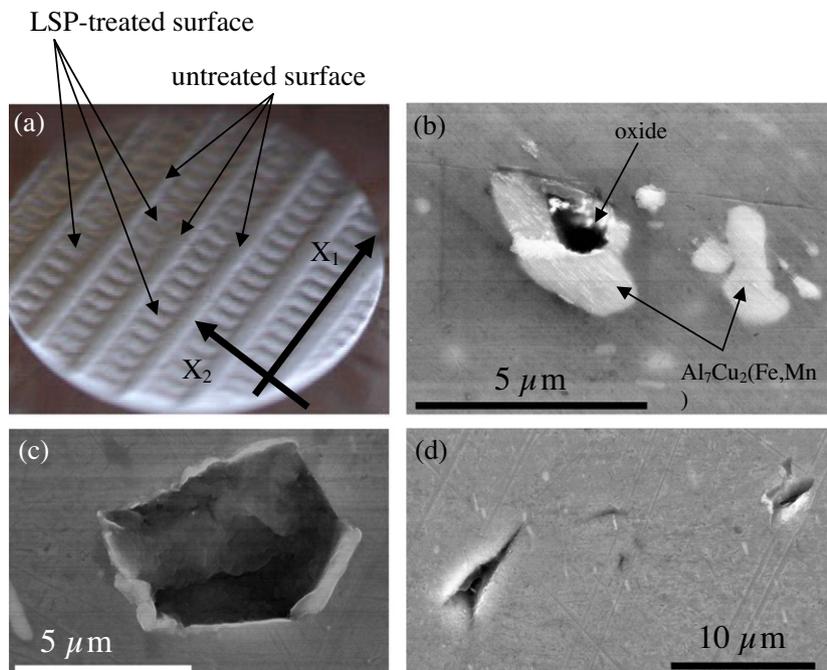
400 and 25  $\mu\text{m}$ . A high-resolution potentiostat (Jaisle IMP83 PCT-BC) was used in these experiments with a current detection limit of approximately 20 fA. At the micro- and macro-scale, polarisation curves were plotted at 16.6  $\text{mV s}^{-1}$  in 0.1 M NaCl (aerated solution) at  $25 \pm 1$   $^{\circ}\text{C}$ . The potential range was from  $-1500$  mV/SCE up to the pitting potential. No prior polarisation was applied to the system.

Local current measurements were performed at OCP in 5 mM NaCl + 50 mM  $\text{Na}_2\text{SO}_4$  (conductivity: 0.783 mS) using an Applicable Electronics SVET. Pt-Ir microelectrodes (MicroProbe Inc.) were black platinised and the sphere diameter after deposition was about 20  $\mu\text{m}$ . The vibration amplitude was 20  $\mu\text{m}$  and the vibration frequency was 600 and 200 Hz in the direction parallel and normal to the surface, respectively. The Y-component of the current was measured. The ASET Software (Science Wares Inc.) converted the potential drop measured by the microelectrode with Ohm's law into a current density value after amplification. The distance between the SVET tip and the sample surface was fixed to 100  $\mu\text{m}$ . SVET measurements were carried out along a line (600  $\mu\text{m}$  in length) across the interface between the LSP-treated surface and the polished surface at different times of immersion. The distance between two measurement points was 2  $\mu\text{m}$ . The displacement of the microelectrode was performed using a motorised and computer-controlled XYZ micromanipulator.

### 2.3. Surface observations and stress measurements at the microscale

A field emission scanning electron microscope (JEOL 6400F) with an integrated electron dispersion spectrometer (FE-SEM/EDS) was used to determine the morphology and chemical composition of constituent particles. AFM images were obtained in air from a Q-Scope 350 microscope (Quesant Instruments Corporation) in the tapping mode (vibration frequency of 174 kHz). Surface roughness values ( $R_a$ ) were calculated from AFM images ( $40 \times 40 \mu\text{m}^2$ ).

Residual stresses ( $\sigma_{11}$  and  $\sigma_{22}$  along the  $X_1$ - and  $X_2$ -axis, respectively) were determined on the polished and LSP-treated surfaces



**Fig. 1.** (a) Optical image of the sample after LSP treatment. (b) FE-SEM/EDS image of constituent particles in AA2050-T8 aluminium alloy after mechanical polishing. FE-SEM/EDS images of surface defects generated by LSP treatment: (c), cavities and (d), microcracks.

using an X-ray  $\theta/\theta$ -type goniometer (D8 Discover with the General Area 2D Detector Diffraction System from Bruker AXS). The X-ray beam was focused using an aperture (diameter of 50  $\mu\text{m}$ ). Experiments were carried out on {3 1 1} planes using the  $\sin^2\psi$  method [20]. The X-ray elastic constant was  $1/2S_2 = 1.9 \times 10^{-5} \text{ MPa}^{-1}$ .

### 3. Results and discussion

The influence of LSP on the microstructure of AA2050-T8 has been investigated in a previous work [21]. FE-SEM/EDS observations revealed the presence of  $\text{Al}_7\text{Cu}_2(\text{Fe},\text{Mn})$  constituent particles, as shown in Fig. 1(b). Their size was in the range of 1–10  $\mu\text{m}$  and their chemical composition was  $73.97 \pm 8.8$  at% Al,  $17.6 \pm 8$  at% Cu,  $5 \pm 2.6$  at% Fe and  $3.3 \pm 2.4$  at% Mn. Oxide particles have also been detected. They contained oxygen ( $22.7 \pm 2$  at%), aluminium ( $67.6 \pm 0.4$  at%), copper ( $2.5 \pm 1$  at%), iron ( $0.6 \pm 0.6$  at%), manganese ( $1.4 \pm 1.3$  at%) and silicon ( $5.1 \pm 1.2$  at%). On the other hand, the matrix was composed of aluminium ( $97.9 \pm 0.09$  at%) and copper ( $2 \pm 0.07$  at%). TEM experiments [22] have revealed that the main hardening phase in AA2050-T8 was  $\text{Al}_2\text{CuLi}$ . It has been shown [21] that LSP treatment has no influence on the size and chemical composition of constituent particles, hardening particles and oxides.

By contrast, LSP affects the residual stress values. Extremely low stress values were registered after mechanical polishing ( $\sigma_{11} = -30 \pm 30 \text{ MPa}$  and  $\sigma_{22} = 15 \pm 25 \text{ MPa}$ ). After LSP treatment, sites containing the pure matrix were found to be under compression ( $\sigma_{11} = -165 \pm 40 \text{ MPa}$  and  $\sigma_{22} = -160 \pm 30 \text{ MPa}$ ).

#### 3.1. Corrosion behaviour of AA2050-T8 in 0.1M NaCl after mechanical polishing

##### 3.1.1. Global polarisation curves

Global polarisation curves in 0.1 M NaCl reveal the presence of a two-stage breakdown mechanism in AA-2050-T8 aluminium alloy after mechanical polishing (Fig. 2(a)). The first stage occurs for applied potentials greater than  $E_{b1} = -785 \text{ mV vs. Ag/AgCl}$ . The second stage is observed for potentials greater than  $E_{b2} = -610 \text{ mV vs. Ag/AgCl}$ . These two stages are associated with a rapid rise in the anodic current density. As previously explained [23,24], they correspond to the dissolution potential of second phase particles.

##### 3.1.2. Corrosion behaviour above the potential $E_{b1}$

To determine the corrosion mechanisms occurring at the first potential  $E_{b1}$ , samples were exposed to the electrolyte at the OCP value for 25 min. The steady state corrosion potential had a value around  $-700 \text{ mV vs. Ag/AgCl}$  (between  $E_{b1}$  and  $E_{b2}$ ). FE-SEM/EDS observations revealed that grain boundaries were not dissolved, but sites containing large constituent particles were preferentially oxidised (Fig. 3).

Some sites were slightly oxidised, as shown in Fig. 3(a). An FE-SEM/EDS analysis revealed that the matrix contains the same quantity of copper as prior to immersion (approximately 2 at%, as shown in Table 1). A small amount of iron (0.13 at%) and manganese (0.66 at%) was detected in the immediate vicinity of constituent particles (Table 1). This suggests that these two elements were dissolved and re-deposited on the surrounding matrix during immersion. Fe and Mn were not detected in the matrix after polishing. Oxygen was preferentially absorbed in constituent particles (18 at% in Table 1) and in the matrix surrounding these particles (8.5 at%). Only 2 at% of oxygen was detected in areas of the matrix far from constituent particles. Increased passivity in areas of the matrix adjacent to constituent particles was then observed.

Other sites were strongly oxidised. An FE-SEM/EDS analysis of the oxide layer formed around constituent particles (Fig. 3(b))

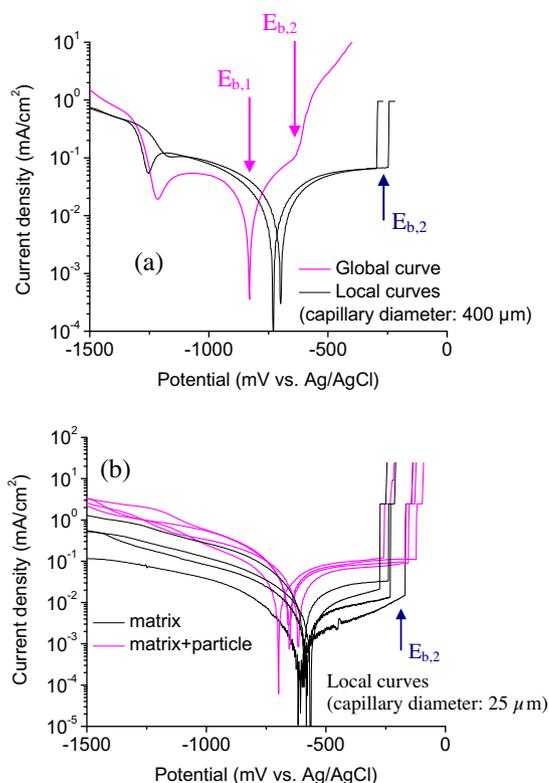


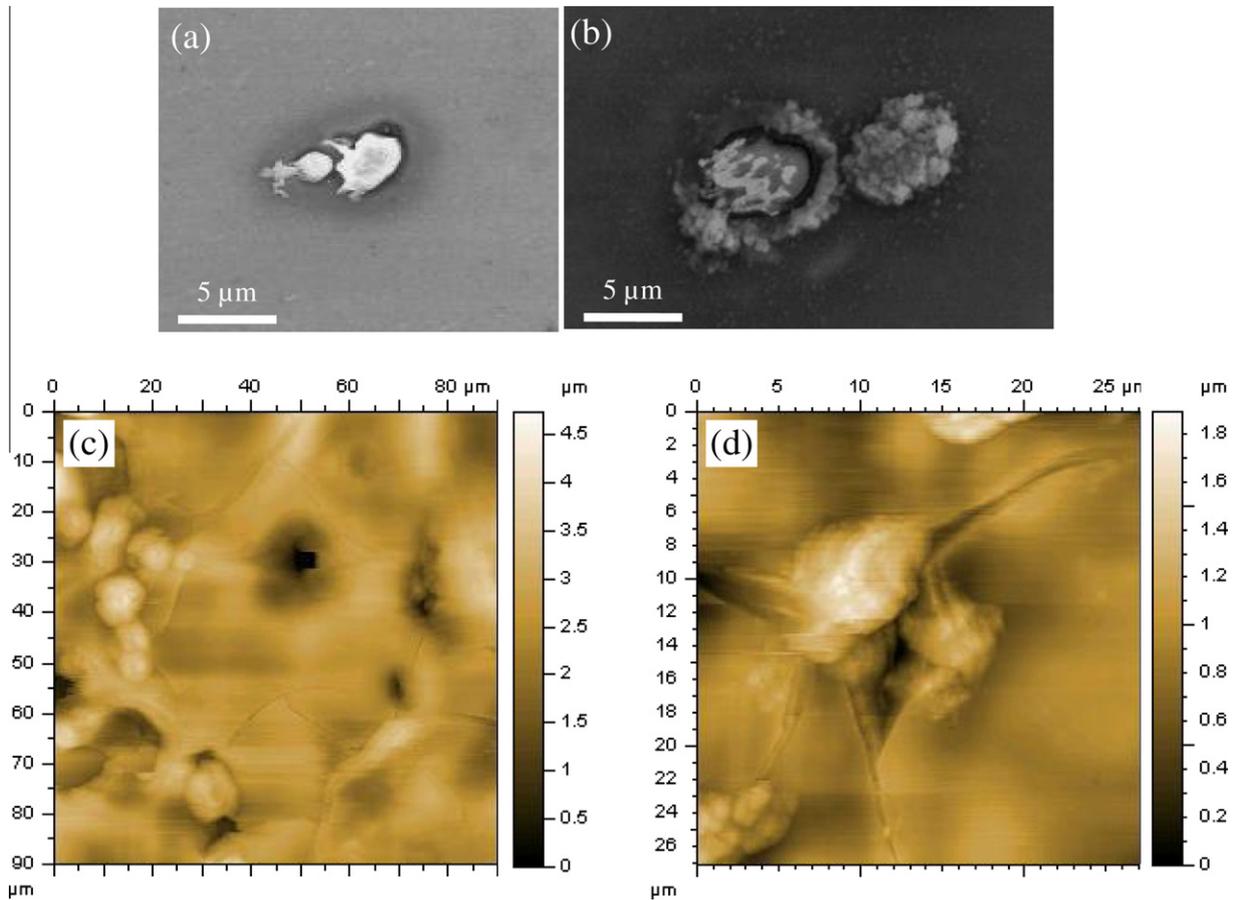
Fig. 2. (a and b) Global and local polarisation curves (16.6 mV/s) for AA2050-T8 aluminium alloy in 0.1 M NaCl at 25°C. Experiments were carried out after mechanical polishing.

revealed the following composition: 48.7 at.% O, 43 at.% Al, 5.05 at.% Cu, 4 at.% Mn, 1.95 at.% Fe and 0.2 at.% Cl. Therefore, aluminium oxide or hydroxide first formed around constituent particles before covering them completely. These surface observations confirm that constituent particles in the 2050-T8 aluminium alloy behave as cathodic sites.

Enrichment in copper of constituent particles (induced by preferential dissolution of Fe and Mn) enhances their cathodic behaviour, and oxygen reduction takes place at their surface. It has been proposed for AA2024-T351 aluminium alloy in 0.1 M  $\text{Na}_2\text{SO}_4 + 0.001 \text{ M NaCl}$  solution [25] that the galvanic coupling between constituent particles (cathodic sites) and the surrounding matrix (anodic site) can explain the increased passive behaviour of the adjacent matrix zone.

##### 3.1.3. Corrosion behaviour above the potential $E_{b2}$

For potentials at  $E_{b2}$  and above, both pitting and intergranular corrosion developed, as shown in Fig. 3(c and d). These figures present AFM images of the sample surface after interruption of the polarisation curve at  $-250 \text{ mV vs. Ag/AgCl}$  (i.e. an applied potential greater than  $E_{b2}$ ). The grain boundaries were attacked and large pits developed within the grains. Two intergranular corrosion mechanisms have been proposed in the literature [26,27], corresponding to the preferential dissolution of the precipitate free zone (PFZ) along grain boundaries or of grain boundary precipitates. In the later case, if the initial grain boundary precipitate is  $\beta'$  or  $\beta$  ( $\text{Mg}_2\text{Al}_3$ ), such a particle could serve as an anodic site and dissolve preferentially. Given that such particles are relatively anodic when compared to aluminium, their dissolution shifts the alloy's potential in the cathodic direction, thereby diminishing the intergranular corrosion driving force. According to reference [27], the intergranular corrosion of 6056-T6 starts at pits. From Fig. 3(d) it seems that this is also the case for 2050-T8.



**Fig. 3.** (a and b) FE-SEM/EDS micrographs of sites with constituent particles after immersion in 0.1 M NaCl at OCP for 25 min. (c and d) AFM images of the specimen surface after the corrosion test corresponding to the global polarisation curve shown in Fig. 2(a).

**Table 1**

Chemical composition of the AA2050-T8 matrix after polishing and after immersion at OCP in 0.1 M NaCl for 25 min (at.%).

	O	Al	Cl	Mn	Fe	Cu
Matrix after polishing	0	97.93	0	0	0	2
Matrix close to constituent particles after immersion at OCP	8.5	88.5	0.045	0.66	0.13	2.2
Matrix far from constituent particles after immersion at OCP	1.7	96.4	0	0.03	0	1.9

Fig. 4 shows that  $E_{b2}$  decreases as the specimen surface area in contact with the electrolyte increases. Pitting potential derived from large scale experiments is usually fixed by the corrosion behaviour of the most active particle whereas the pitting potential derived from an experiment on a single particle is usually determined by the weakest zone of the particle.

#### 3.1.4. Local polarisation curves

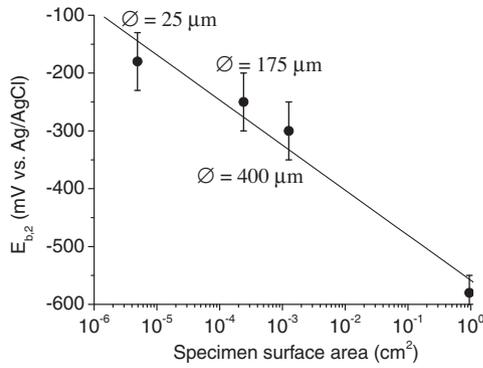
Only the second stage (i.e. potentials above  $E_{b2}$ ) is visible in the local polarisation curves (Fig. 2(a and b)). A similar result was obtained using three capillary diameters. It has been shown [24,28] that the number of stages in a polarisation curve is related to the surface area of the sample. Only a single stage has been detected for very small surface areas in contact with the electrolyte. It has been suggested that constituent particles do not all react at the same time. As a result of the low number of constituent particles within a small surface area, the anodic current relating to their dissolution is extremely low and the associated stage (i.e. potentials above  $E_{b1}$ ) is not visible. Any differences observed between

macro- and micro-measurements were therefore associated with the volume fraction of constituent particles in the material.

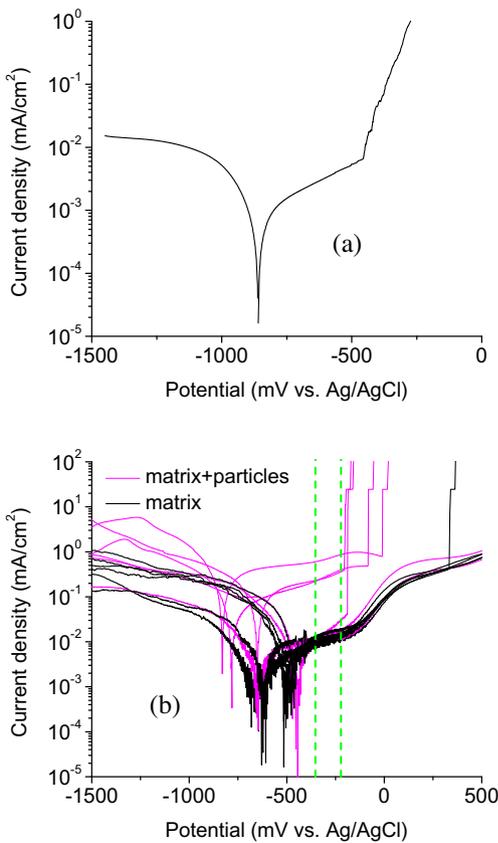
### 3.2. Corrosion behaviour of AA2050-T8 aluminium alloy in 0.1 M NaCl after LSP

#### 3.2.1. Global polarisation curves

To determine the influence of LSP on the electrochemical behaviour of AA2050-T8, measurements were carried out at sites of the pure matrix that were free of any visible defects. The global polarisation curve in 0.1 M NaCl is shown in Fig. 5(a). The two stages are observed on this curve. The more active potential ( $E_{b1}$  at approximately  $-850$  mV vs. AgCl/AgCl) is once again associated with the oxidation of sites containing constituent particles. This was confirmed by surface observations that were made after corrosion tests under potentiostatic control (at  $-630$  mV vs. Ag/AgCl, between  $E_{b1}$  and  $E_{b2}$ ), as shown in Fig. 6(a and b). The surface morphology obtained was very close to that found in the absence of LSP treatment. For potentials greater than the second breakdown



**Fig. 4.** Evolution of the pitting potential  $E_{b2}$  of AA2050-T8 vs. the specimen surface area exposed to the electrolyte. Data obtained both from global and local measurements.

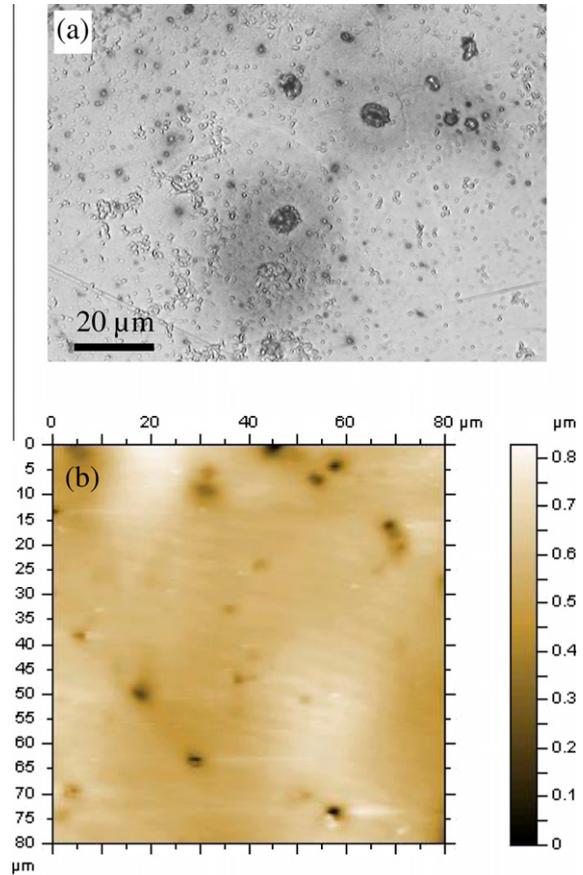


**Fig. 5.** (a) Global polarisation curve (16.6 mV/s) for AA2050-T8 aluminium alloy in 0.1 M NaCl at 25 °C. The experiment was performed on the LSP-treated surface. (b) Local polarisation curves obtained under the same experimental conditions using 25 μm diameter capillaries.

potential ( $E_{b2}$  at  $-450$  mV vs. Ag/AgCl), numerous pits are observed in grains, as shown in Fig. 6(b). These pits initiated at sites containing constituent particles. In contrast with the previous case, no intergranular corrosion developed in the matrix. Therefore, it can be concluded that LSP treatment has beneficial effects on the corrosion resistance of the matrix.

### 3.2.2. Local polarisation curves

Fig. 5(b) shows the local polarisation curves obtained in 0.1 M NaCl, at sites containing the pure matrix and at those containing the matrix with a cavity. These sites exhibit a large passive range.



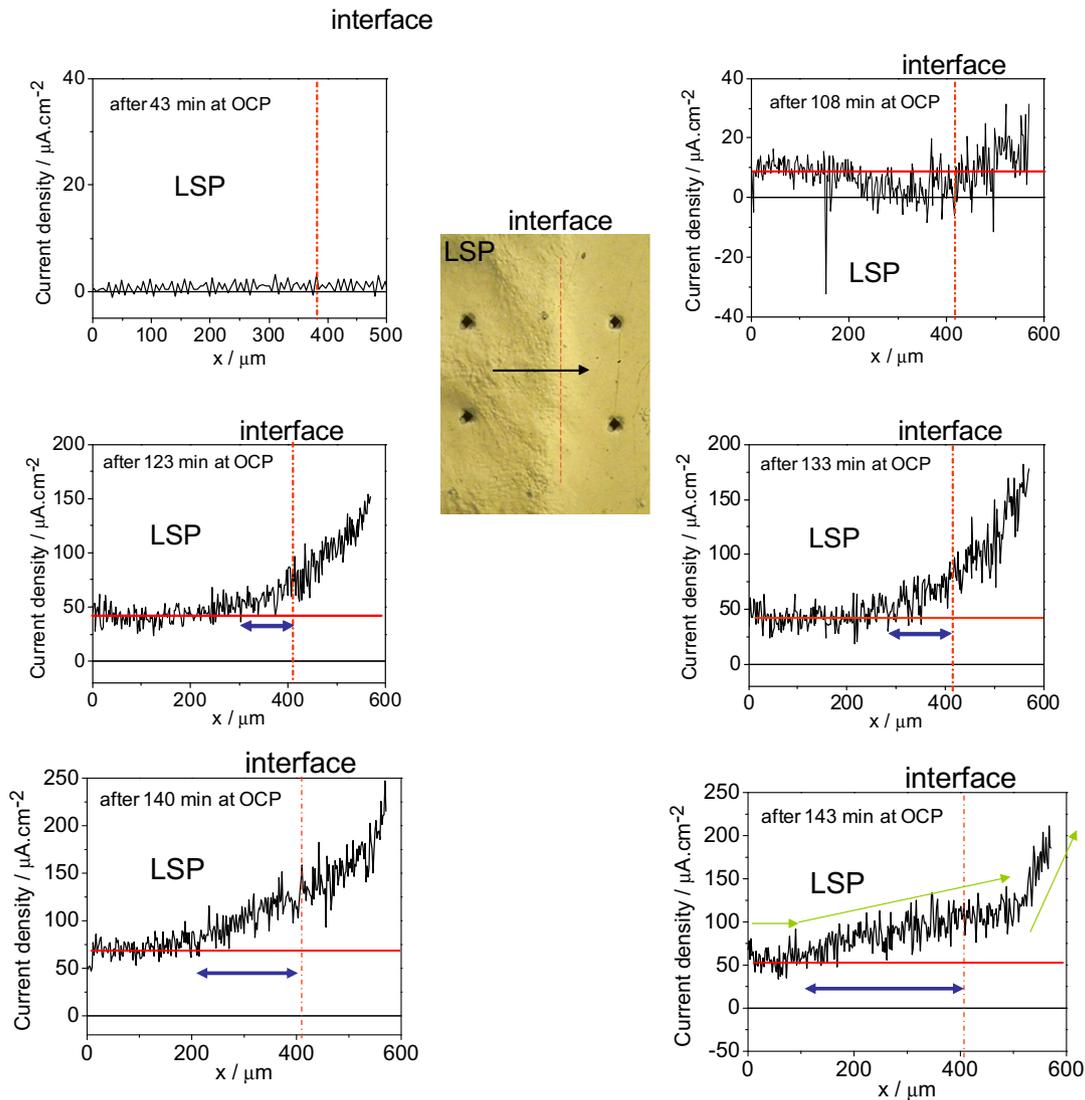
**Fig. 6.** (a) Optical images of the specimen surface after a potentiostatic test at  $-630$  mV vs. Ag/AgCl for 20 min. (b) AFM image of the specimen surface after the corrosion test corresponding to the global polarisation curve shown in Fig. 5(a).

The current density was low ( $0.01 \text{ mA cm}^{-2}$ , as opposed to  $0.05\text{--}0.1 \text{ mA cm}^{-2}$  after polishing), and no pitting potential could be identified. Therefore, LSP has beneficial effects on the electrochemical behaviour and corrosion resistance of the pure matrix. This is mainly attributed to the formation of a surface under compression.

After LSP treatment, scattering was observed in the local polarisation curves of sites containing constituent particles, as shown in Fig. 5(b). The current density was between  $0.25$  and  $0.65 \text{ mA cm}^{-2}$  in the anodic domain, and higher than  $1 \text{ mA cm}^{-2}$  in the cathodic region. Pitting potentials were slightly more anodic than they were after polishing ( $-50$  mV vs. Ag/AgCl, as opposed to  $-100$  mV vs. Ag/AgCl). Therefore, LSP has little influence on the behaviour of sites with constituent particles. Microstructural investigations of the surface were performed after the local measurements. However, these investigations were not helpful to interpret which type of corrosion occurs and which type of corrosion process is responsible for the obtained curves. XRD measurements showed that the surface is under compression after LSP treatment. In the future, thermal-mechanical simulation based on the use of the finite element method will be used to predict mechanical effects induced in the material treated by LSP. Surface stresses may play an important role in corrosion processes after LSP. Local polarization curves will be analysed considering these mechanical effects.

### 3.3. Corrosion behaviour of AA2050-T8 aluminium alloy in low chloride concentration solution using SVET

SVET was then used to map the current distribution across the interface between the polished and the LSP-treated surfaces at



**Fig. 7.** Evolution of the local current density vs. the immersion time. SVET measurements were carried out in 5 mM NaCl + 50 mM Na<sub>2</sub>SO<sub>4</sub>, across the interface between the polished and the LSP-treated surfaces.

OCP. A low chloride concentration solution was chosen because SVET measurements are usually performed in low conductivity media to improve the spatial resolution. Conversely, measurements with microcapillaries need high conductivity media due to the very low measured current (small analysed surface area).

The evolution of the normal current density vs. the immersion time is shown in Fig. 7. For 43 min of immersion, no current was measured. The surface was passive on both sides of the interface. After 108 min of immersion, a current density of 20  $\mu\text{A cm}^{-2}$  was evidenced on the polished surface far from the interface, versus only around 10  $\mu\text{A cm}^{-2}$  on the LSP-treated surface. This indicates that the polished surface is more sensitive to corrosion than the LSP-treated one. It can be mentioned that negative currents were measured at some points on both sides of the interface. As it was discussed in Section 3.1.2, cathodic reactions take place at particles (size in the range of 1–10  $\mu\text{m}$ ). The spatial resolution of the SVET does not allow the detection of a current on small particles. Therefore, it was difficult to detect negative currents.

For 123 min of immersion, the current density increased by a factor of 7.5 on the polished surface and only by a factor of 5 on the LSP-treated surface, confirming that the polished surface is more prone to corrosion than the LSP-treated surface. For 123 min of immersion, the current density remains more-or-less

constant on the LSP-treated surface (around 50  $\mu\text{A cm}^{-2}$ ), whereas it increases up to 200  $\mu\text{A cm}^{-2}$  on the polished surface. These results indicate that LSP has beneficial effects on corrosion processes. Indeed, preferential dissolution of the LSP-treated surface near the interface was observed for 123 min of immersion.

#### 4. Conclusions

The corrosion behaviour of AA2050-T8 aluminium alloy was studied after polishing and after laser shock processing (LSP) treatment using the electrochemical microcell technique and the scanning vibrating electrode technique (SVET). The following conclusions can be drawn:

- (1) LSP has no influence on the microstructure of AA2050-T8 aluminium alloy. This alloy contains Al<sub>7</sub>Cu<sub>2</sub>(Fe,Mn) constituent particles. The main hardening phase was Al<sub>2</sub>CuLi.
- (2) Extremely low stress values were registered after mechanical polishing. By contrast, the matrix was found to be under compression after LSP.
- (3) After polishing, both pitting corrosion at constituent particles and intergranular corrosion were observed. By contrast, no intergranular corrosion developed in the matrix after LSP.

- (4) After LSP, microcell measurements revealed that pitting potential slightly increased (of about 50 mV vs. Ag/AgCl) in sites containing constituent particles. By contrast, the matrix remains always passive. This was attributed to the presence of compressive stresses.

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