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Topography and wettability characterization of surfaces manufactured by SLM and treated by chemical etching

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ABSTRACT

Selective Laser Melting process represents an interesting opportunity in the biomedical field to fabricate customized implants. However, the surface roughness of components obtained through additive manufacturing is a major limitation and affects the surface wettability. In the present work, chemical etching is adopted to deal with such an issue. To do so, the effects of chemical etching parameters (such as immersion time and composition of the solution) on the surface roughness, weight loss and wettability is analyzed. Different samples (obtained through different printing orientations) are considered. The tests show that the roughness and the wetting of the surfaces are improved thanks to chemical etching. As a major result, the most influencing parameters on surface wetting are two: the roughness and the material properties (which vary with samples depth).

KEYWORDS

Additive manufacturing; topography; wettability; selective laser melting; biomedical implants; chemical etching

1. Introduction

The additive manufacturing (AM) process is an innovative technology that makes it possible to manufacture complex 3D geometries with minimization of material waste [1]. One of the main advantages of the process is great freedom in terms of the geometry of the components and the possibility to use different materials for different regions during the process [2]. Besides, AM seems promising for challenging issues where weight minimization is a priority factor. Among the existing AM processes, those based on the fusion of a powder bed are the most developed ones. An example is the Selective Laser Melting (SLM) process [3]. SLM process consists in the addition of melting layers of a metallic powder bed, using a laser as power source. The interaction between the metallic powder and the laser creates the “selective melting” of specific parts of the layer according to the configuration of the slice. The same operation is repeated on all the slices composing the object, allowing the layer-by-layer manufacturing process. Such a process can be used for various applications, such as rapid prototyping, the manufacturing of spare parts, the creation of molds [4], etc. In particular, this process could let obtain very promising applications in the biomedical technologies. Indeed, it can be used to manufacture customized implants which can be adapted to the specific morphology of the patient through the use of radiological data. Moreover, the designer can act on the mesoscopic structure of the implant to optimize

some required properties (e.g. stiffness, strength, weight, etc.).

Among the several aspects to be analyzed to completely characterize the properties and the physical behavior of components issued from SLM process, the surface roughness is of paramount importance especially for specific applications (e.g. aeronautical, biomedical, etc.). The experimental data on surface roughness of components issued from more classic industrial processes show a high variability, going from less than $1\ \mu\text{m}$ for machining [5, 6] to more than $100\ \mu\text{m}$ for casting [7]. In the case of the SLM, the surface exhibits a distinctive roughness, with a mean value around $15\ \mu\text{m}$ [8–11]. This resulting roughness has various origins: (i) the staircase effect due to the orientation α of the surface during the manufacturing [12–14]; (ii) the texture created by melting bed during the process [15]; (iii) the sintered powder particles which are trapped in the neighboring of the melting bed area. Lhuissier et al. [16] defined (iii) as type I surface defect, and (i) and (ii) as type II surface defects. Type I surface defects can generate several issues, such as mechanical defects, or tolerance errors [17].

Indeed, the surface roughness obtained from the process could affect the surface interactions with its environment. In the literature and depending on the field of applications, it has been shown that the surface roughness could boost or reduce the performance of the surface [18, 19]. For instance, it has been proven that the adhesive bond with a liquid state phase is higher on rough surfaces [20]. However, beyond an optimal value of the roughness, the strength of adhesive

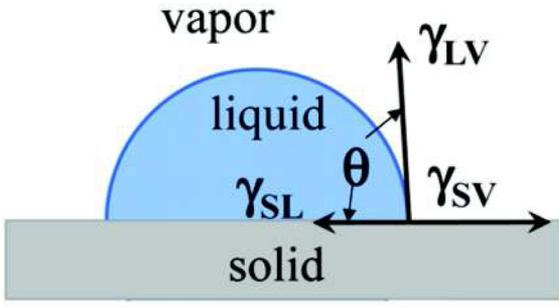


Figure 1. Schematic of the equilibrium of the phases from the study of Law et al. [21].

bonds starts decreasing. Speaking of the adhesive bond, the thermodynamic adhesion has to be considered when talking about interactions. From a thermodynamic point of view, adhesion can be defined as a thermodynamic work necessary to keep the equilibrium state of liquid, solid, and vapor phases (see Figure 1). This definition takes the form of the Dupre equation [21]:

$$W_a = \gamma_{LV} + \gamma_{SV} - \gamma_{SL}. \quad (1)$$

with W_a the work of adhesion, γ_{LV} the liquid/gas interfacial energy, γ_{SV} the solid/gas interfacial energy, and γ_{SL} the solid/liquid interfacial energy.

Very often, interfacial bonding is evaluated through the wettability of the surface [22]. Thus, wetting data may be used to estimate the thermodynamic adhesion of the surface [23, 24]. Wettability is commonly characterized through the sessile drop method, that consists in the static measurement of a contact angle θ (see Figure 1) that arises from the equilibrium between solid, liquid, and gas phases [25]. When using water as a liquid phase, a surface is defined as hydrophilic for $\theta < 90^\circ$ and hydrophobic for $\theta \geq 90^\circ$ [21, 26]. It must be pointed out that this definition is not absolute. Using this method, we are able to calculate the surface free energy γ_{SL} , using the Young-Dupre relation [27]:

$$\gamma_{SV} = \gamma_{SL} + \gamma_{LV} \cos \theta. \quad (2)$$

Typical values of wettability measured on SLM surfaces containing powder particles seem to show a hydrophobicity of the surface with contact angle $\theta \sim 100^\circ$ for Ti-6Al-4V (TA6V) [28], reflecting a low surface energy, thus a low thermodynamic adhesion. However, several fields of applications promote the hydrophilic surfaces, with contact angles inferior to 90° . More precisely, in the case of machined TA6V flat surface, $\theta \in [70^\circ, 90^\circ]$ [29]. Therefore, in the case of SLM components, the hydrophobic surface properties can be assigned to the presence of sintered particles trapped into the surface.

In order to solve this issue, surface treatment can be considered. Among the existing surface treatments, the chemical etching seems to be the most adapted to the SLM process. Indeed, chemical etching is compatible with the 3D complexity of typical parts issued from the SLM process, and it is relatively easy to execute [16]. This typical chemical etching process consumes the Ti element on the surface. Then, the treatment enables to smooth the surface, allowing to decrease the surface roughness and increasing the surface

wettability. Indeed, Zahran et al. [30] shows that for an etching solution of 10% HF on commercially pure polished Ti, a decrease of contact angle from $\theta \sim 60^\circ$ to $\theta \sim 40^\circ$ is observed. Therefore, contrary to several studies on flat surfaces treated using chemical etching to increase roughness, and so, to improve wetting [31], on SLM-based surfaces, the primary goal of the acid etching is to remove the sintered particles from the surface to decrease the roughness, and so to try to maximize wettability. Thus, the knowledge of the chemical etching kinetic on the topology and the wetting on SLM-based surfaces is of fundamental importance.

In this paper, the aim is to improve the roughness and the wettability of TA6V surfaces obtained by SLM. Although TA6V alloy, issued from classical manufacturing processes, has been investigated many times, the originality of our approach consists in analyzing the chemical etching process in order to improve the surface conditions of SLM-based surfaces in terms of roughness and wettability. In particular, the variation of roughness due to the printing orientation α , with respect to the printing plate, will be considered. The study will focus on the application to the biomedical field. More precisely, medical devices (such as implants or prostheses) must present a higher sensitivity to wetting, with a high hydrophilicity of surfaces in order to boost the cellular adhesion and proliferation on the surface [32, 33]. For instance, Vogler et al. [34] have shown that a better initial cell adhesion and a larger contact area of the cell is registered on wettable surfaces compared to hydrophobic materials. On the other hand, Li et al. [35] have shown an increase of cell detachment on the hydrophobic surface. Besides, the roughness is also a strictly related parameter to consider. Indeed, Wennerberg et al. [36] show that a moderate roughness, between 1 and $2 \mu\text{m}$, gives an enhanced bone response and promotes better cell adhesion, motility, and shape, which are essential for efficient early osseointegration. Soro et al. [37] evaluated the impact of etching on TA6V lattices specimens for biological applications and have shown the complex nature of the local roughness variation as a function of the edging time and local surface orientation with respect to the building direction, fully supporting the need for the detail analyses proposed in the present work. Thus, to attain the main goal of the present work, the chemical etching parameters to obtain high surface energy with moderate hydrophilicity ($\theta < 90^\circ$) and low roughness value (average roughness belonging to $[1 \mu\text{m} - 2 \mu\text{m}]$) is determined. To this purpose, the first part of the study focuses on the characterization of the SLM surface and its different properties. Then, the parameters of the chemical etching and their influence on the surface roughness and wettability are analyzed. To finish, an analysis of the chemical etching depending on α is made to evaluate its effectiveness. Some general remarks and conclusions end the paper.

2. Materials and methods

2.1. Sample preparation and chemical etching protocol

Disks of titanium alloy TA6V were fabricated with a thickness of 3 mm and a diameter of 11.5 mm . Those disks were produced from TA6V powder, using the AM machine

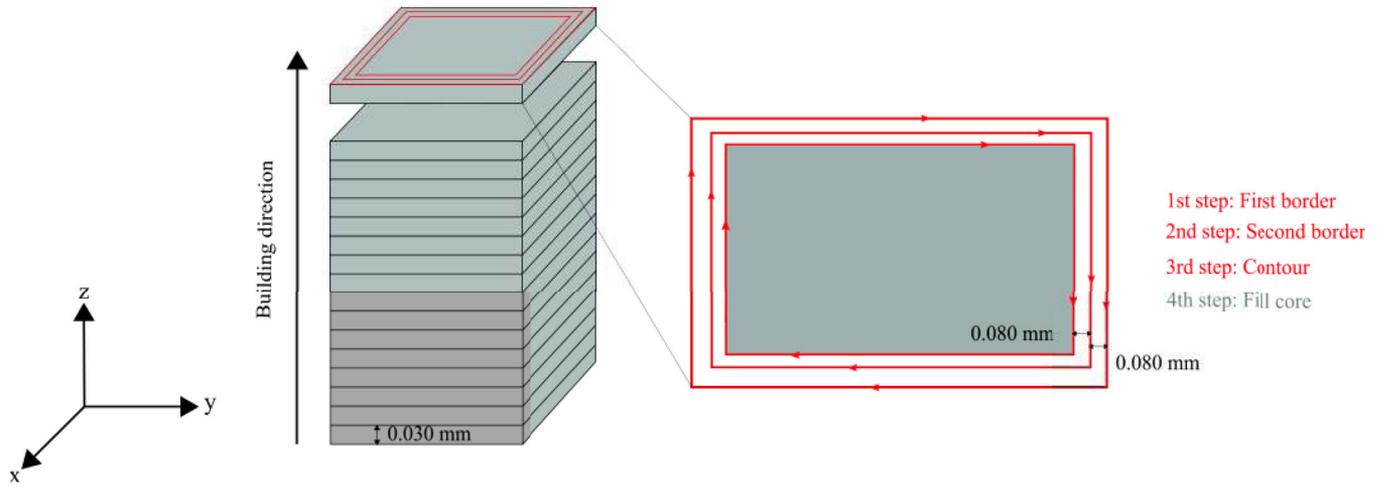


Figure 2. Scan strategy of the SLM process. Each layer is represented by a rectangular parallelepiped.

Table 1. Values of SLM nominal parameters for the borders and contour, and the core.

Borders and Contour parameters		Core parameters	
Power P (W)	150	Power P (W)	275
Scan speed V (mm/s)	450	Scan speed V (mm/s)	1100
Hatch distance h (mm)	0.08	Hatch distance h (mm)	0.12
Layer thickness e (mm)	0.03	Layer thickness e (mm)	0.03
Energy density E (J/mm ³)	138.9	Energy density E (J/mm ³)	69.4

SLM²⁸⁰ (2.0 of SLM SOLUTIONS GROUP AG). Each layer of the AM process was 30 μm thick and was generated using a two-step laser path: first, the external perimeter of the layer (defined by two borders and one contour in red in Figure 2) was fused; then, the internal core was filled (defined by the gray region in Figure 2).

The borders and contours were generated using an energy density of 139 J/mm³ whereas the internal core was generated using an energy density of 69 J/mm³. Each energy density was defined by a nominal value of laser power P, scan speed V, hatching space h, and layer thickness e, whose values are defined in Table 1. After fabrication, the support structures were removed and the samples were cleaned using ethanol, then distilled water within an ultrasonic tank in order to remove the unsintered powder.

To analyze the chemical etching process, three protocols were defined in this study:

Protocol n° 1.

A set of disks oriented at $\alpha = 90^\circ$ (plane (x, y) being the printing plane, see Figure 3) were fabricated, then chemically etched in a mixtures of hydrofluoric and nitric acids. Three different acid baths were tested on this set of samples. For each bath, 5 different times of immersion were tested: 5 min, 10 min, 15 min, 20 min, and 30 min. Acid composition of the batches are summarized in Table 2.

Protocol n° 2.

A set of disks still oriented at $\alpha = 90^\circ$ were produced, then polished with a set of abrasive papers (chronological sequence: P80 \rightarrow P100 \rightarrow P400 \rightarrow P800 \rightarrow P1200 \rightarrow P2400). A chemical etching in a solution composed of 3%

HF and 10% HNO₃ was then applied with 5 different times of immersion: 5 min, 10 min, 15 min, 20 min, and 30 min.

Protocol n° 3.

Four sets of specimens were produced by SLM, each set being defined by a different inclination α of the plane of the disk with respect to the (x,y) printing plane (Figure 3): 0°, 30°, 60°, and 90°. For each set, a chemical etching in a solution composed of 3% HF and 10% HNO₃ was then applied for 30 minutes.

For each protocol, each chemical etching was realized within an ultrasonic tank in order to obtain a homogeneous etching on all the surface. After the chemical etching, the samples were cleaned using first ethanol, then distilled water within an ultrasonic tank in order to remove all traces of the etching solution. The mass of each sample was measured before and after chemical polishing.

2.2. Surface topography analysis

3D surface roughness data of each sample were acquired using an optical profilometer AltiSurf 500 (AltiMet, France) in chromatic confocal mode. The area of 1 mm \times 1 mm was measured with an in-plane resolution of 5 μm and a vertical resolution of 10 nm. The surface measurements were made following a regular grid mapping of the surface and a grid step size of 5 μm . The roughness was measured orthogonally to the inclination plane. The 3D surface roughness measurements were made on two representatives samples of each surface type (raw SLM, Protocols n° 1, 2, 3) and repeated 3 times randomly on the surface. The optical profilometer provided data on the topographic parameters: arithmetic mean roughness R_a , root mean square roughness R_q , the maximum height R_z , skewness R_{sk} , Kurtosis parameter R_{ku} . The definition of all the parameters can be found in the study of Gadelmawla et al. [38]. Besides, we also calculated the developed area ratio S_w , defined as the ratio between the area of the real surface and the area of the projected surface. In addition, the morphology of each material surface was examined under a white light microscope VHX-5000 (Keyence, Itasa, USA). The morphology was evaluated using

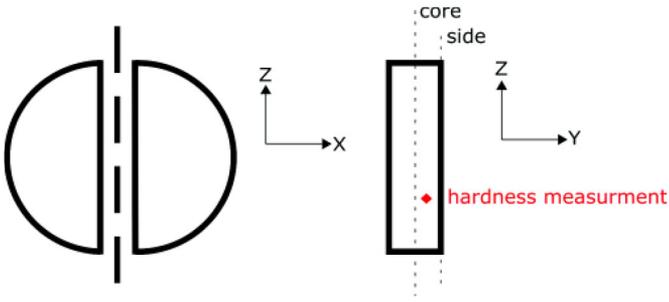


Figure 3. Sample cut in the Y-Z plane for hardness measurement.

Table 2. Composition of the hydrofluoric and nitric acids baths for the chemical etching.

n°	HF (%)	HNO ₃ (%)
1	3	10
2	3	20
3	10	10

1600 × 1200 pixel images and 3D reconstruction of the surface using the software included with the system. Moreover, a reference group of untreated SLM surface with an inclination of $\alpha = 90^\circ$ was examined under a scanning electron microscope (Zeiss EVO50 HD, Germany).

2.3. Wettability and surface energy measurement

In the case of a free ideal solid surface the contact angle correspond to the definition of Young's contact angle. However, when dealing with real surfaces presenting micro and nano-roughness the notion of apparent contact angle θ_{app} must be used. The apparent contact angle is defined as the angle between the apparent surface and the liquid-fluid interface tangent, see [23, 24, 39, 40]. In the present work, apparent contact angle θ_{app} and surface energy measurements were made using the contact angle system DSA 30 (Kruss, Germany). Nine measurements on three samples of each surface type (raw SLM, Protocols n° 1, 2, 3) were performed using 2 media (distilled water and diiodomethane). During the tests, 5 μ L of each media were placed on the surface, then analyzed using the software provided with the DSA 30 system. The surface energy was determined from the apparent contact angle θ_{app} by post-processing, using the Owen-Wendt method [25]. Each media was characterized by a polar component σ_L^p and a dispersive component σ_L^d . From the contact angle measurement and by regression, we were able to determine the polar (σ_S^p) and dispersive (σ_S^d) component of the surface, using the following equation:

$$\frac{\sigma_L(1 + \cos \theta_{app})}{2\sqrt{\sigma_L^d}} = \sqrt{\sigma_S^p} \sqrt{\frac{\sigma_L^p}{\sigma_L^d}} + \sqrt{\sigma_S^d}. \quad (3)$$

Then, the free energy surface (equal to $\sigma_S^p + \sigma_S^d$) was calculated. In the results section (Sec. 3), all θ_{app} are measured using water.

2.4. Hardness measurement

The surface stiffness was evaluated using micro-indentation hardness tests. One representative sample from the reference group of untreated SLM surface with inclination $\alpha = 90^\circ$ was chosen, and cut in the local Y-Z plane to obtain the section of the sample (see Figure 3). Hardness measurements from the external border to the internal core were taken. Hardness measurements were performed with a 500 g load using a Vicker hardness tester. A series of 10 tests were conducted along evenly-spaced 150 μ m interval from the external border to the internal core and the process was repeated three times. In order to analyze the relation between hardness and wetting, a set of disks were produced by with an inclination of $\alpha = 90^\circ$, then machined along the y direction to remove an external layer of thickness t . Six values of t were set: 0.2 mm, 0.3 mm, 0.4 mm, 0.5 mm, 0.7 mm, and 0.9 mm. 5 samples were analyzed for each value of t .

2.5. Statistical analysis

Data were analyzed with a Kruskal Wallis test, followed by a post-hoc Spearman test. The Kruskal Wallis test analyses the null hypothesis and is defined by a p-value p which will be considered as significant if $p < 0.05$. The Spearman test analyses the correlation between two sets of data and is defined by a correlation value c and a p-value p (also in this case, p will be considered as significant if $p < 0.05$).

3. Results

3.1. Characterization of SLM surfaces

Figure 4 depicts the morphologies of surfaces obtained for 4 different orientations α of the specimen with respect to the printing plane (i.e. $\alpha = 0^\circ, 30^\circ, 60^\circ$, and 90°). Table 3 presents the values of Ra , Rq , Rz , Rsk , Rku and S_w parameters relative to the surfaces at the 4 different orientations. At $\alpha = 0^\circ$, the surface presents a wavy morphology created by the cooling of the melting bed. An increase of roughness is observed when α passes from 0° to 30° , with an increase of average roughness Ra from 9.35 μ m (± 1.44) to 15.33 μ m (± 1.07) (see Table 3). Rq , Rz , and S_w follow the same evolution. The negative value of Rsk at $\alpha = 30^\circ$ points out the predominance of peaks on the surface. We can assume that the increase of roughness from $\alpha = 0^\circ$ to 30° results from the combining effect of type II defects, via the staircase effect (the AM of a tilted surface takes place by the superposition of shifted layers) which is more pronounced with α , and type I defects (sintered particles) on the surface (see Figure 4b). The Kurtosis value (Rku) is larger than 3 for the two orientations. This confirms a predominance of sharp peak distribution on $\alpha = 0^\circ$ and 30° surfaces, accentuated by the type I defects at $\alpha = 30^\circ$.

When α changes from 30° to 60° , the morphology evidences an increase of density of type I defects (sintered particles) with an increase of S_w from 2.32 (± 0.09) to 2.81 (± 0.07) (see Table 3). A decrease of Ra is simultaneously observed whereas no significant differences are found for Rq

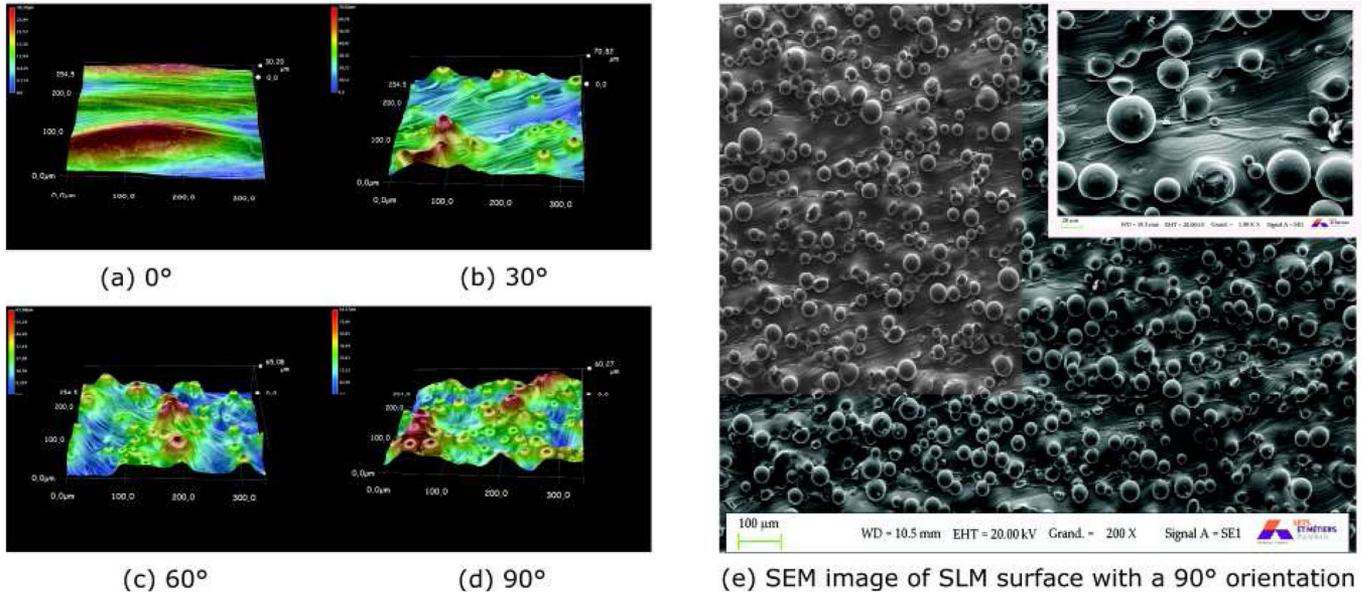


Figure 4. 3D morphologies of the surface orientations. (a:d) gives the surface topographies for $\alpha=0^\circ$, 30° , 60° , and 90° . (e) gives a SEM image of the surface at $\alpha=90^\circ$.

Table 3. Values of roughness parameters depending on α .

	Ra (μm)	Rq (μm)	Rz (μm)	Rsk (μm)	Rku (μm)	S_w
0°	9.35 (± 1.44)	11.89 (± 1.56)	64.29 (± 12.15)	-0.21 (± 0.18)	4.02 (± 1.08)	1.376 (± 0.2)
30°	15.33 (± 1.07)	20.25 (± 1.11)	133.13 (± 4.42)	0.88 (± 0.21)	4.8 (± 0.56)	2.316 (± 0.09)
60°	13.77 (± 0.79)	18.19 (± 1.17)	134.93 (± 11.72)	0.96 (± 0.24)	5.22 (± 1.26)	2.81 (± 0.073)
90°	15.75 (± 1.15)	19.9 (± 1.2)	130.8 (± 9.6)	0.56 (± 0.4)	3.71 (± 0.11)	3.17 (± 0.17)

and Rz . Therefore, we can assume that the increase of roughness with the density of type I defects is counterbalanced by the decrease of roughness due to the mitigation of the staircase effect for higher α . Finally, when α changes from 60° to 90° , the roughness of the sample surface keeps reducing. Therefore, this seems to follow the considerations still made for $\alpha=30^\circ$ and 60° . The roughness deriving from the type II defects (staircase effect) becomes negligible compared the roughness due to the type I defects when α increases. An increase of Ra and S_w is measured, whereas Rz and Rq do not vary considerably. A decrease of Rku value from $\alpha=60^\circ$ to 90° is measured. Rsk is still positive, underlining the predominance of peaks created by the particles.

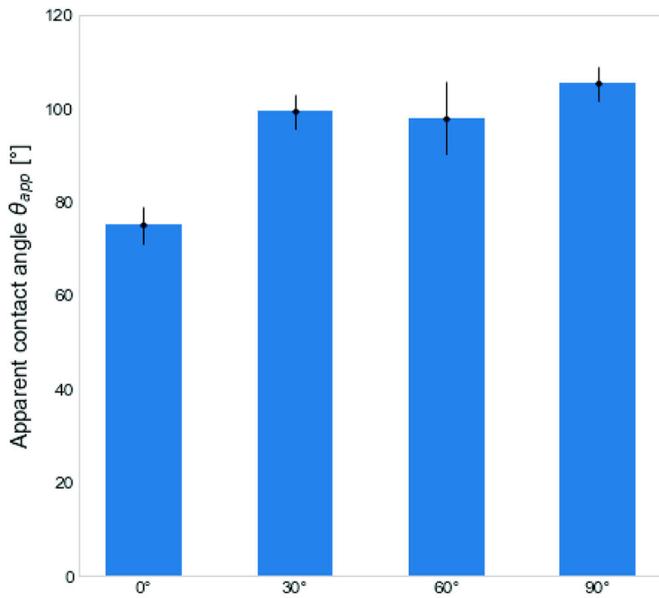
Concerning the wettability of raw SLM samples, surfaces obtained are mostly hydrophobic (whereby $\theta_{app} > 90^\circ$) (Figure 5). Concerning samples fabricated at $\alpha=0^\circ$, the surface presents a θ_{app} of $75^\circ (\pm 4.1)$ for a surface energy of $60 \text{ mJ/m}^2 (\pm 4.1)$. Thus, this surface can be defined as moderately hydrophilic.

Increasing the fabrication angle α from 0° to 30° results in an increase of θ_{app} up to $99.3^\circ (\pm 3.7)$, and a relative decrease of surface energy down to $40.5 \text{ mJ/m}^2 (\pm 1.9)$. For specimens fabricated at an inclination greater than 30° , no significant differences are found in terms of θ_{app} and surface energy ($p=0.85$, Kruskal-Wallis method). No correlations are found between Ra and the θ_{app} or the surface energy ($p=0.2$, Spearman method). However, the variation of θ_{app} when α passes from $\alpha=0^\circ$ to 30° seems to match the emergence of type I defects. One assumption can be that the presence of type I defects could cause the hydrophobicity of the surface.

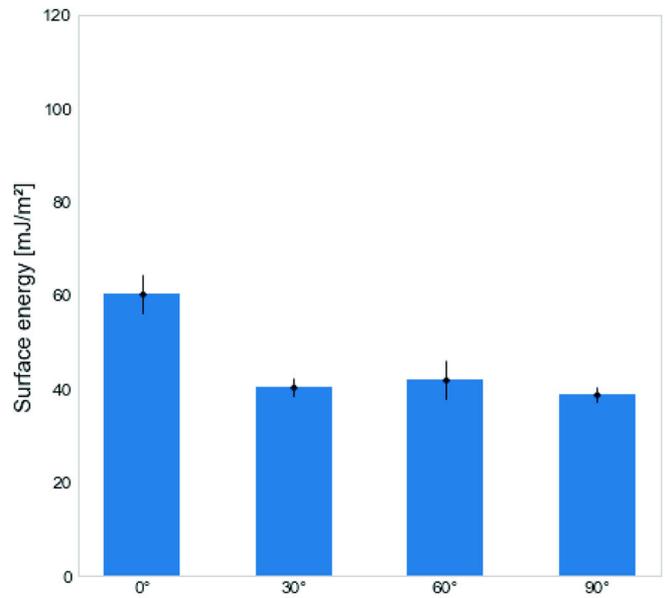
3.2. Characterization of SLM TA6V material

To characterize the evolution of material properties inside SLM-based components, an analysis of microhardness has been realized on $\alpha=90^\circ$ oriented samples according to hardness measurement tests described in Section 2.4. Figure 6 shows an increase of hardness with the depth, from $341 \text{ HV} (\pm 1.5)$ at 0.2 mm to $368 \text{ HV} (\pm 2)$ at 0.5 mm depth from the outer surface of the sample. From 0.5 mm to 0.9 mm , no significant differences are found ($p=0.6$, Kruskal-Wallis method), proving a stabilization of the material properties. Those results highlight the presence of a gradient of material properties along the thickness of SLM-based components.

Figure 7 presents the results of θ_{app} and surface energies for different in depth values obtained through the same machining process (see Section 2.4). At an in depth of 0.2 mm , we measure a θ_{app} of $86.4^\circ (\pm 4.2)$ and a surface energy of $49.3 \text{ mJ/m}^2 (\pm 3.45)$. No significant differences are found between θ_{app} at in depth of 0.2 , 0.3 , and 0.4 mm ($p=0.78$, Kruskal-Wallis method). A significant decrease of θ_{app} is measured when passing from an in depth of 0.4 mm to 0.5 mm ($p=0.009$, Kruskal-Wallis method), whose θ_{app} varies from $85.8^\circ (\pm 7.2)$ to $69.6^\circ (\pm 7.8)$. This is reflected by an increase of surface energy from $50.2 \text{ mJ/m}^2 (\pm 5.1)$ to $66.4 \text{ mJ/m}^2 (\pm 8.6)$. Then, a stabilization of the value of θ_{app} and surface energy is measured, with no significant differences between the different in depth surface of 0.5 , 0.7 , and 0.9 mm ($p=0.3$, Kruskal-Wallis method). This result matches the trend of the gradient of material.



(a) Values of θ_{app}



(b) Values of surface energies

Figure 5. Values of θ_{app} and surface energy depending on α .

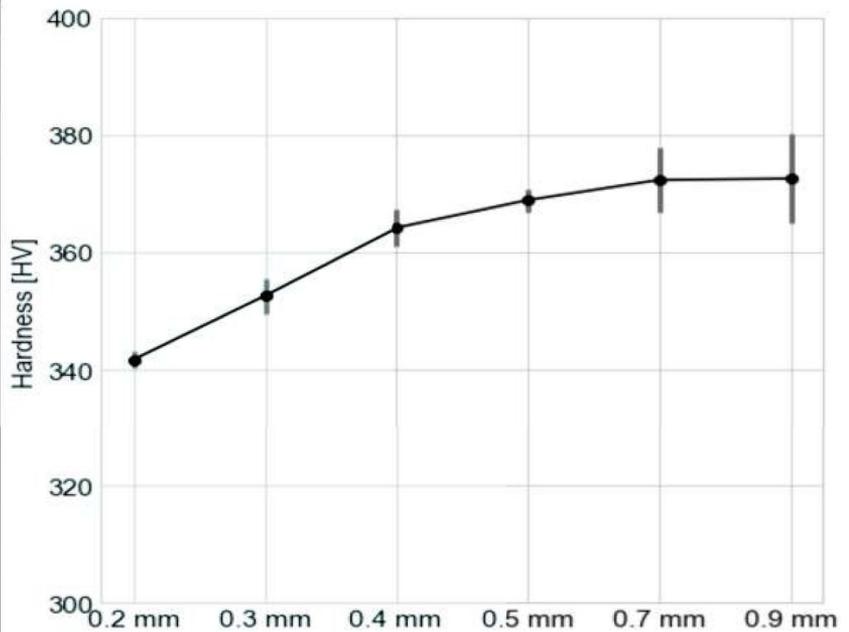


Figure 6. Value of hardness as function of the distance from the outer surface.

3.3. Chemical etching analysis

Figure 8 shows the morphology of SLM-based surfaces treated with chemical etching using protocol n° 1. In terms of morphology, Figure 8 shows the elimination of type I defects at $t = 5$ min, for all the chemical etching compositions. Besides, a decrease of surface amplitude with the increase of immersion time is also observed for those 3 compositions with however significant differences between them. 3% HF/10% HNO₃ and 3% HF/20% HNO₃ treatments keep the traces of the SLM process and a resulting anisotropy of the surface even after 30 min of immersion, whereas 10% HF/10% HNO₃ removes all process traces on the surface.

Table 4 gives the value of the roughness parameters for the different compositions and immersion times. For a bath composition of 3% HF/10% HNO₃, a decrease of R_a is measured, from 14.5 μm (± 2.6) at 5 min to 6.8 μm (± 0.7) at 30 min. R_q and R_z follow the same evolution. R_{sk} takes positive values during the etching, with thus a predominance of peaks, except for 30 min where local porosities (i.e. local cavities) start to appear (reflected by $R_{sk} < 0$). A decrease of R_{ku} is also observed during the etching, from 5 min to 30 min. However, those values stay higher than or equal to 3. This can be read as a decrease of the sharp peaks in the surface in favor of more rounded peaks. In terms of

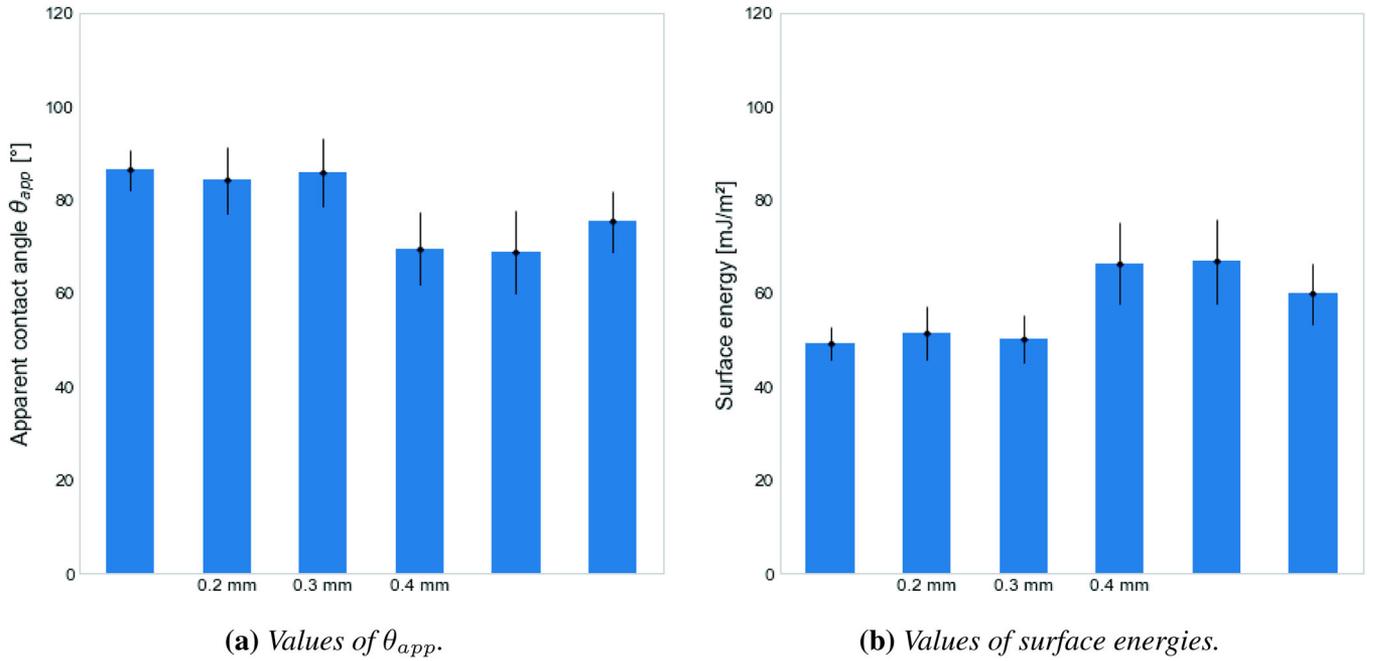


Figure 7. θ_{app} and surface energies as function of the distance to the surface.

Table 4. Values of roughness parameters for the different chemical etching compositions and durations.

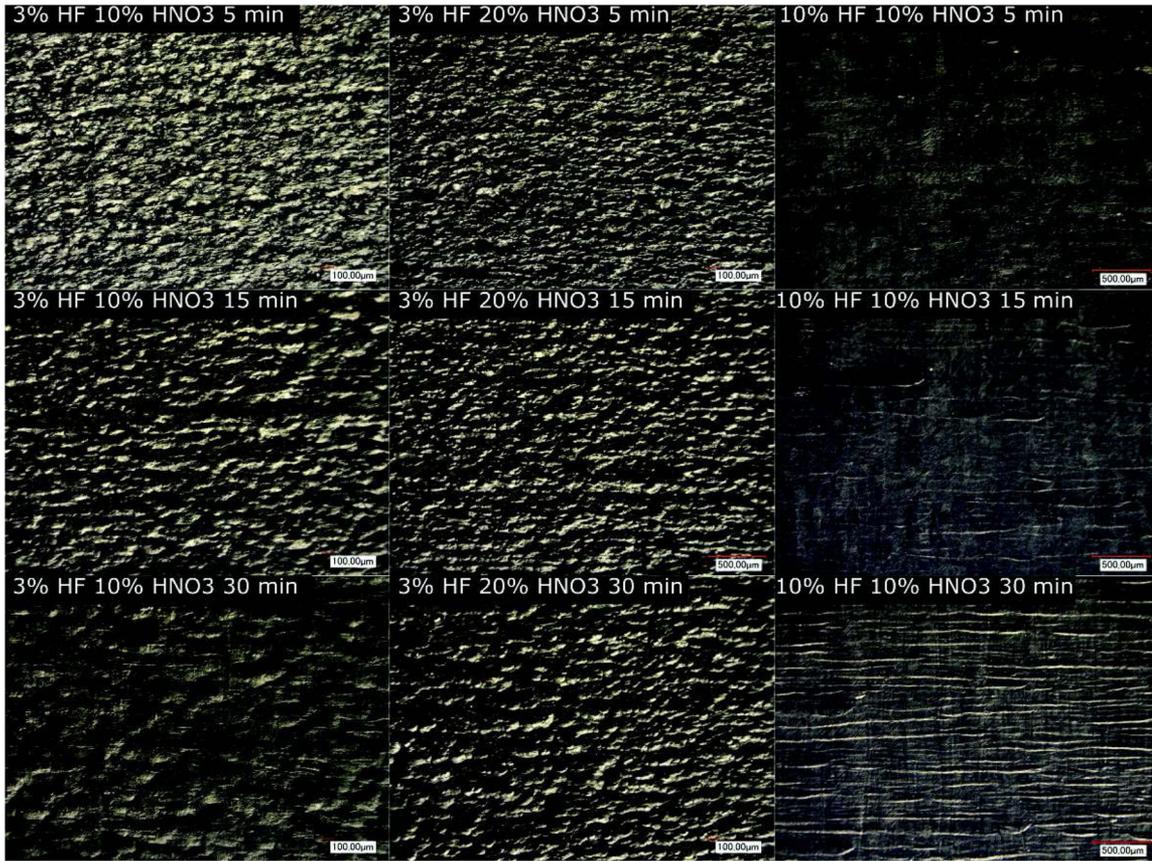
Chemical polishing composition	time (min)	Ra (μm)	Rq (μm)	Rz (μm)	Rsk	Rku	S_w
3% HF / 10% HNO_3	5	14.5 (± 2.55)	16.7 (± 2.2)	101.3 (± 13.6)	0.47 (± 0.13)	3.75 (± 0.09)	2.02 (± 0.07)
	10	12.8 (± 1.94)	16.3 (± 2.7)	95.5 (± 18.7)	0.64 (± 0.07)	3.63 (± 0.4)	1.71 (± 0.2)
	15	9.7 (± 1.9)	12.2 (± 1.7)	74.1 (± 8.6)	0.27 (± 0.17)	3.26 (± 0.4)	1.41 (± 0.08)
	20	8.4 (± 2.23)	10.7 (± 2.8)	50.8 (± 6.2)	0.32 (± 0.18)	3.15 (± 0.31)	1.34 (± 0.07)
	30	6.8 (± 0.66)	8.4 (± 0.6)	40.7 (± 4.6)	-0.03 (± 0.28)	3 (± 0.8)	1.26 (± 0.09)
3% HF / 20% HNO_3	5	12.1 (± 2.8)	16.2 (± 3)	111 (± 13.3)	0.85 (± 0.75)	5.79 (± 1.4)	1.99 (± 0.21)
	10	10.7 (± 1.6)	13.6 (± 2.1)	89.2 (± 16.1)	0.23 (± 0.18)	3.8 (± 0.5)	1.75 (± 0.18)
	15	10.3 (± 1.3)	12.9 (± 1.5)	80.1 (± 11.1)	0.13 (± 0.18)	3.2 (± 0.36)	1.5 (± 0.11)
	20	7.6 (± 1.4)	9.9 (± 1.2)	59.6 (± 13.7)	-0.23 (± 0.44)	3.1 (± 0.4)	1.47 (± 0.11)
	30	9.2 (± 1.6)	11.8 (± 1.2)	58.9 (± 19.9)	-0.18 (± 0.2)	3.41 (± 0.97)	1.45 (± 0.15)
10% HF / 10% HNO_3	5	4.8 (± 0.1)	6 (± 0.14)	27.3 (± 1.6)	0.085 (± 0.06)	2.65 (± 0.07)	1.25 (± 0.03)
	10	2.9 (± 0.8)	3.6 (± 0.3)	12.7 (± 2.6)	-0.042 (± 0.06)	2.85 (± 0.07)	1.2 (± 0.04)
	15	2.5 (± 0.5)	3.2 (± 0.7)	13.6 (± 2.9)	-0.33 (± 0.91)	2.92 (± 0.3)	1.14 (± 0.19)
	20	3 (± 1.2)	3.8 (± 1.5)	15.9 (± 3.2)	-0.05 (± 0.15)	3.63 (± 1)	1.09 (± 0.04)
	30	2.2 (± 0.3)	2.8 (± 0.6)	15 (± 2.8)	-0.18 (± 0.2)	3.6 (± 0.9)	1.03 (± 0.06)

developed surface S_w , a decrease of S_w is observed with the increase of immersion time, from 2.02 (± 0.07) at 5 min to 1.26 (± 0.09) at 30 min. A bath composition of 3% HF/20% HNO_3 results in a similar trend. A decrease of roughness is registered from 5 min to 20 min, with the decrease of Ra from 12.1 μm (± 2.9) down to 7.6 μm (± 1.37). We observe then an increase of Ra at 30 min with a value of 9.2 μm (± 1.6). The same trend can be seen for Rq although no significant difference is found for Rz . In terms of surface texture, we notice that for a time lower or equal to 15 min, the surface is mostly composed of peaks, with however a decrease of Rsk . Rsk becomes negative at $t=20$ min, revealing the creation of local porosities on the surface. Rku follows the same trend as Ra , but stays always higher than 3 during all the etching times, showing that sharp asperities remain on the surface. In terms of S_w , we observe a decrease down to 20 min at 1.47 (± 0.11). No significant differences are found for S_w values between 20 and 30 min. For a bath composition of 10% HF/10% HNO_3 , a significant decrease of roughness is found during the first 5 min, with the decrease of Ra down to 4.8 μm (± 0.1). We observe then a slower decrease

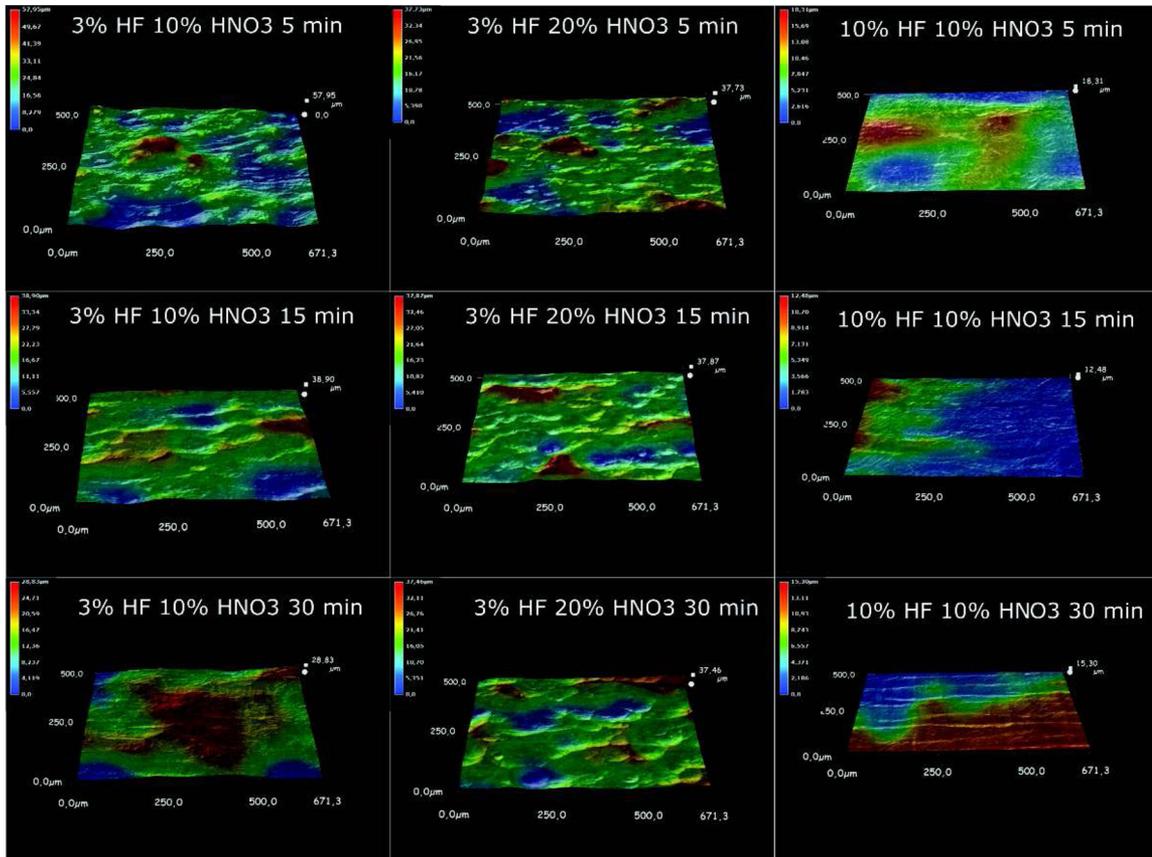
of Ra down to 2.2 μm (± 0.3) at 30 min. Rq and S_w follow the same trend, although no significant differences are found for Rz between 15 and 30 min. In terms of surface texture, Rsk keeps being negative during the process, showing an increase of local porosities. This is supported by the increase of Rku , until 20 min where it become higher than 3. This can be interpreted as an increase of asperities with the increase of local porosities.

Figure 9 shows the chemical etching kinetics for different acid bath compositions and times of immersion, via the weight loss and the thickness loss of the surface. The thickness loss parameter corresponds to half the thickness difference between the untreated and the treated sample. We took the half to represent the loss of material on one side of the sample (see Figure 3).

The kinetic of 3% HF/10% HNO_3 shows a loss of 0.4%/min of sample mass during the first 5 minutes, corresponding to a loss of 20 $\mu\text{m}/\text{min}$ of the surface. After 5 min, we observe a decrease of the reaction speed, with a weight loss of 0.6%/min from 10 min to 30 min, corresponding to a thickness loss of 9 $\mu\text{m}/\text{min}$, reaching a weight loss of 16% (\pm

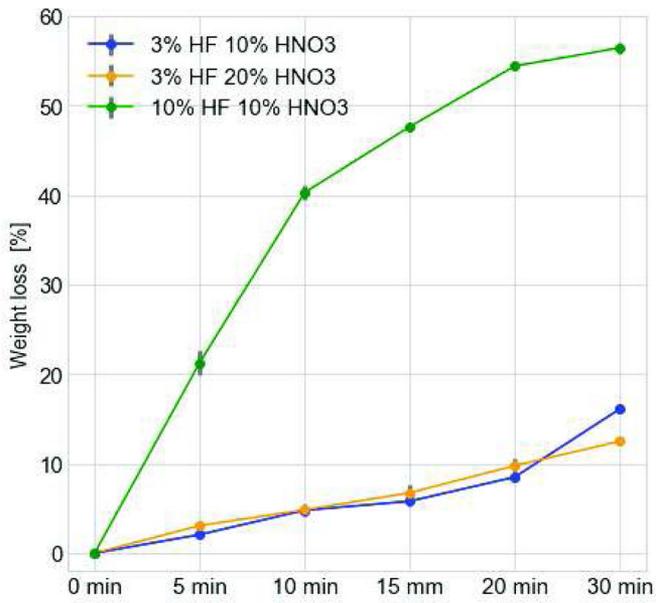


(a) *White light microscopy.*

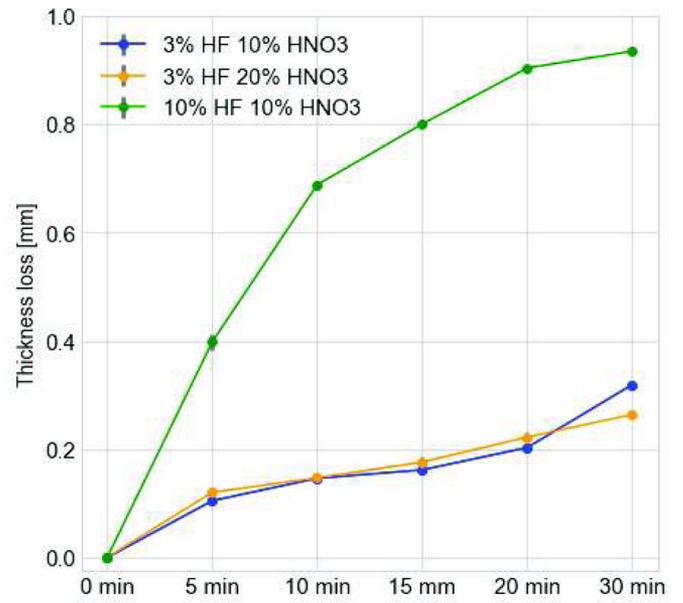


(b) *3D morphologies.*

Figure 8. Chemical etching surface for different compositions and durations. Each column represents a composition (from left to right: 3% HF / 10% HNO₃, 3% HF / 20% HNO₃, 10% HF / 10% HNO₃). Each line represents a time of immersion (from top to bottom: 5, 15, and 30 minutes).



(a) Values of weight loss.



(b) Values of thickness loss.

Figure 9. Weight loss and thickness loss parameters for the different acid bath compositions and durations (green: 10% HF / 10% HNO₃, orange: 3% HF / 20% HNO₃, blue: 3% HF / 10% HNO₃).

0.2) and a thickness loss of $319 \mu\text{m} (\pm 3)$ at 30 min. The kinetic of 3% HF/20% HNO₃ follows a similar evolution, with however some differences. A weight loss of 0.6%/min during the first 5 minutes is observed, corresponding to a thickness loss of $24 \mu\text{m}/\text{min}$. Then, from 10 min to 30 min, an average weight loss of 0.4%/min is measured, corresponding to a thickness loss of $6 \mu\text{m}/\text{min}$. At 30 min, we reach a weight loss of 12% (± 0.3) and a thickness loss of $264 \mu\text{m} (\pm 3)$. The kinetic of 10% HF/10% HNO₃ is the most severe, with a weight loss of 4%/min during the first 5 minutes, corresponding to a thickness loss of $80 \mu\text{m}/\text{min}$. When passing from 10 to 30 min, we observe then a weight loss of 0.8%/min, corresponding to a thickness loss of $12 \mu\text{m}/\text{min}$. At 30 min, we obtain a weight loss of 56% (± 0.6) and a thickness loss of $934 \mu\text{m} (\pm 6)$.

Figure 10 gives the wetting properties of the etched surfaces, as function of the chemical etching compositions and times of immersion.

Surfaces treated with 3% HF/10% HNO₃ acid bath show a significant decrease of the θ_{app} during the first 15 minutes, until $48.1^\circ (\pm 10.4)$. Then, an increase of θ_{app} is registered until 20 min for a value of $75.8^\circ (\pm 5.5)$, followed by a stabilization of the θ_{app} at 30 min for $76.3^\circ (\pm 7.8)$. Simultaneously, the results show an increase of surface energy up to $90.6 \text{ mJ}/\text{m}^2 (\pm 11)$ at 15 min, then a decrease down to $61 \text{ mJ}/\text{m}^2 (\pm 5.2)$ at 20 min. No significant differences are found between the surface energy at 20 and 30 min ($p=0.89$, Kruskal-Wallis method). For an acid bath of 3% HF/20% HNO₃, a significant decrease of the θ_{app} is observed during the first 10 minutes, until $44.6^\circ (\pm 9.1)$, followed by an increase of θ_{app} at 15 min for a value of $77^\circ (\pm 6.7)$. No significant difference is found between 15, 20, and 30 min ($p=0.5$, Kruskal-Wallis method). Simultaneously, we observe an increase of surface energy up to $95 \text{ mJ}/\text{m}^2 (\pm$

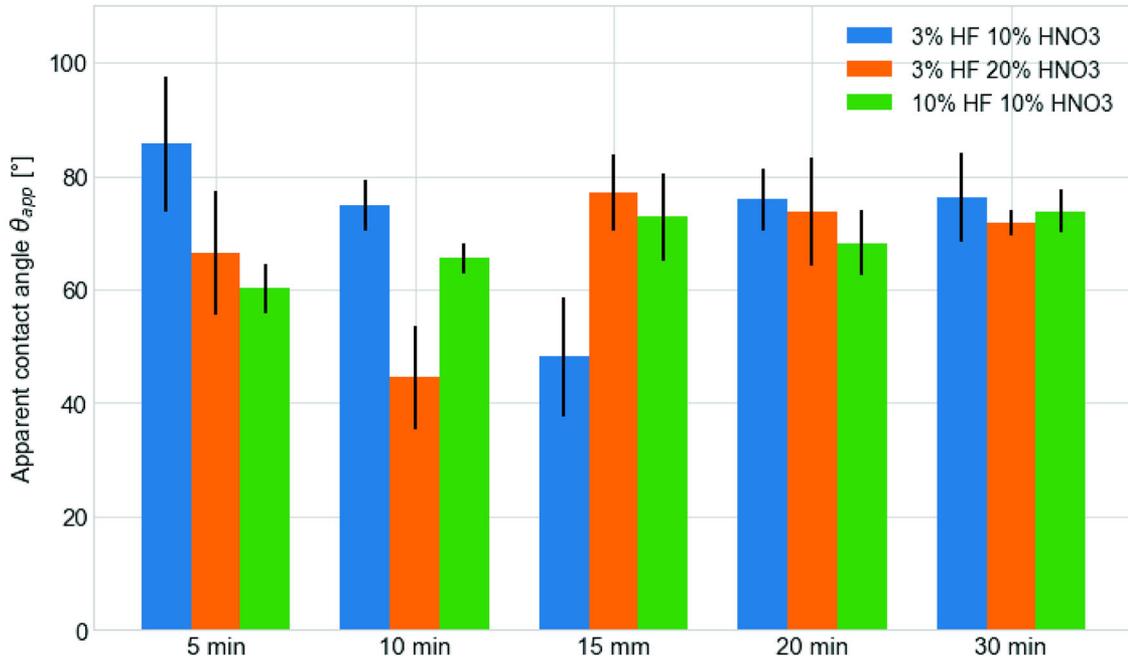
9.4) at 10 min, then a decrease down to $60.6 \text{ mJ}/\text{m}^2 (\pm 6.4)$ at 20 min. No significant differences are found between the surface energies at 15, 20, and 30 min ($p=0.15$, Kruskal-Wallis method). The wetting behavior of surfaces treated with 10% HF/10% HNO₃ acid bath shows a significant decrease of the θ_{app} during the first 5 minutes, until $60.3^\circ (\pm 4.3)$, followed by an increase at 10 min for a value of $65.5^\circ (\pm 2.8)$. No significant difference is found between 10, 15, 20, and 30 min ($p=0.4$, Kruskal-Wallis method). Simultaneously, we have an increase of surface energy to $73.9 \text{ mJ}/\text{m}^2 (\pm 5.2)$ at 5 min, then a decrease at $67 \text{ mJ}/\text{m}^2 (\pm 3.2)$ at 10 min. No significant differences are found between the surface energy at 10, 15, 20, and 30 min ($p=0.4$, Kruskal-Wallis method).

3.4. Topography VS chemical etching

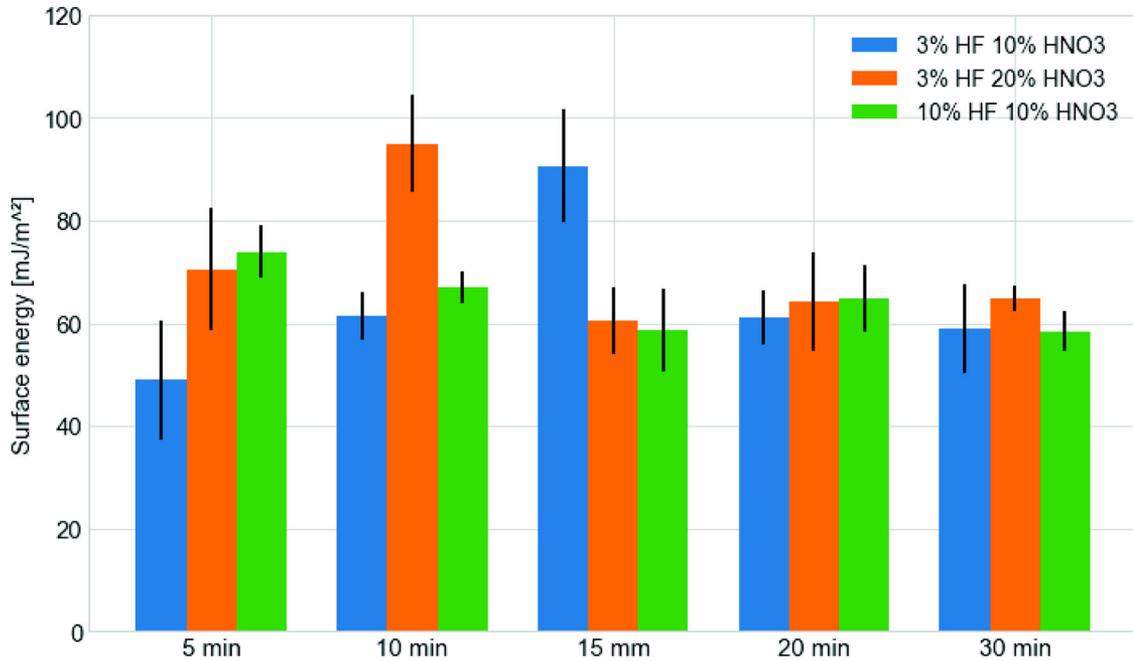
3.4.1. Chemical etching on polished surface

To assess the effect of the surface original topography on the chemical etching, an analysis on chemical etching on polished surface has been made. To do this, we chose the chemical etching 3% HF/10% HNO₃ for 5 different immersion times.

Figure 11 and Table 5 show the morphologies and the roughness parameters of SLM-based surfaces polished and treated with chemical etching using protocol n° 2. Figure 11a shows a reference sample after 15 min of chemical etching. Taking into account the data from Table 5, the results show that the surface is relatively flat, with however a slight increase of roughness over the time of immersion. R_a varies from $0.9 \mu\text{m} (\pm 0.2)$ at the initial state to $1.61 \mu\text{m} (\pm 0.22)$ at 20 min. Then, a decrease at $1.45 \mu\text{m} (\pm 0.1)$ is observed at 30 min. The same trend is observed for R_q and S_w . Also, R_z follows the same evolution for the first 15 minutes.



(a) Values of θ_{app}



(b) Values of surface energies

Figure 10. θ_{app} and surface energies for the different acid bath compositions and durations (green: 10% HF / 10% HNO₃, orange: 3% HF / 20% HNO₃, blue: 3% HF / 10% HNO₃).

However, no significant differences are found for Rz between 20 and 30 min. In the same time, Rsk becomes negative after 20 min, revealing the generation of local porosities on flat surface. During all the treatment, an increase of Rku is observed from 3.2 (± 0.8) to 6.2 (± 2.39), revealing an increase of sharp asperities distribution when increasing the immersion time.

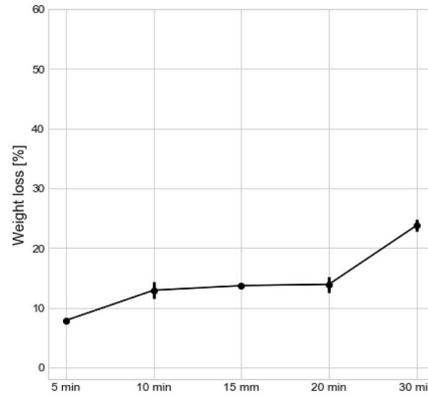
In terms of chemical etching kinetics, the evolutions of weight loss and thickness loss presented in Figure 11b and 11c show that we obtain a similar behavior of chemical

etching on raw SLM surface (see Figure 9 for 3% HF/10% HNO₃), except during the first 5 minutes. We observe an average weight loss of 0.6%/min from 10 to 30 min, corresponding to a thickness loss of 8 μm /min. Thus, we can assume that only the very external surface roughness affects the chemical etching kinetic (with type I defects for instance).

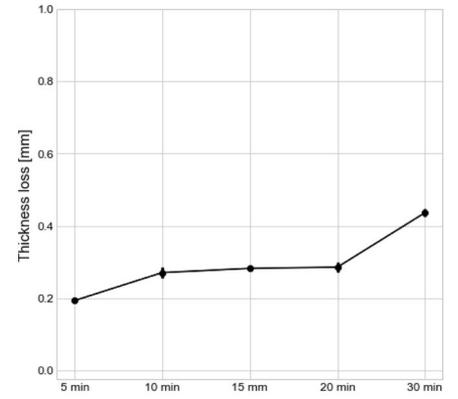
In terms of wettability of the surfaces, Figure 12 presents θ_{app} and surface energies for the different times of immersion on polished surfaces. The initial polished surface is



(a) Surface morphology for a chemical etching duration of 15 minutes.



(b) Values of weight loss.



(c) Values of thickness loss.

Figure 11. Results of chemical etching on polished surface for an acid bath composition of 3% HF / 10% HNO₃.

Table 5. Values of roughness parameters for the different etching times.

	Ra (μm)	Rq (μm)	Rz (μm)	Rsk (μm)	Rku (μm)	S_w
untreated	0.9 (\pm 0.2)	1.5 (\pm 0.1)	10.1 (\pm 1.1)	0.05 (\pm 0.01)	3.2 (\pm 0.8)	1.01 (\pm 0.01)
5 min	1.12 (\pm 0.11)	1.8 (\pm 0.38)	10.47 (\pm 1.25)	0.02 (\pm 0.01)	4.07 (\pm 0.05)	1.03 (\pm 0.01)
10 min	1.37 (\pm 0.14)	2.2 (\pm 0.49)	12.38 (\pm 1.19)	0.12 (\pm 0.14)	4.25 (\pm 0.78)	1.04 (\pm 0.01)
15 min	1.3 (\pm 0.13)	2.11 (\pm 0.49)	11.03 (\pm 1.51)	0.25 (\pm 0.14)	5.25 (\pm 2.33)	1.06 (\pm 0.02)
20 min	1.61 (\pm 0.22)	2.6 (\pm 0.64)	16 (\pm 1.58)	-0.06 (\pm 0.06)	4.95 (\pm 1.63)	1.12 (\pm 0.02)
30 min	1.45 (\pm 0.1)	2.37 (\pm 0.69)	18.57 (\pm 3.2)	-1.3 (\pm 0.24)	6.2 (\pm 2.39)	1.09 (\pm 0.02)

defined by a θ_{app} of $78.6^\circ (\pm 6)$, corresponding to a surface energy of $52.7 \text{ mJ/m}^2 (\pm 6.2)$. No significant differences are found in the θ_{app} during the first 15 minutes ($p=0.13$, Kruskal-Wallis method), and the same comments hold for the surface energy. We observe then a decrease of θ_{app} at 20 min with a value of $70^\circ (\pm 6.1)$, corresponding to an increase of surface energy up to $64.2 \text{ mJ/m}^2 (\pm 6.8)$. Then, at 30 min, θ_{app} increases again up to $75.4^\circ (\pm 7.2)$, corresponding to a decrease of surface energy down to $57.5 \text{ mJ/m}^2 (\pm 7.4)$.

3.4.2. Chemical etching on specimens at different orientations α

Figure 13 shows the morphology of four SLM-based surfaces fabricated at different inclination (i.e. $\alpha=0^\circ$, 30° , 60° and 90°) and treated with chemical etching using protocol $n^\circ 3$. The 3D morphologies point out the absence of type I defects on all the surface with a decrease of surface amplitude.

The corresponding roughness parameters are depicted in Table 6. Ra decreases with the orientation when α passed from 0° to 60° , whose values vary from $7.88 \mu\text{m} (\pm 0.56)$ to $4.79 \mu\text{m} (\pm 0.21)$. Then, Ra increases at $\alpha=90^\circ$ with the value of $6.51 \mu\text{m} (\pm 1.39)$. The same trend can be observed for Rq and Rz . Rsk assumes negative values for $\alpha=0^\circ$, 30° , and 90° whereas at 60° , $Rsk \sim 0$. In addition, Rku is almost equal to 3 for all values of α . Compared to the raw surfaces (see Table 3), the application of protocol $n^\circ 3$ generates a decrease of sharp peak distribution and an increase of local porosities on all surfaces. However, differences are registered for S_w , where we observe an increase from $1.3 (\pm 0.02)$ at $\alpha=0^\circ$ up to $1.35 (\pm 0.03)$ at 30° , then a decrease down to

$1.23 (\pm 0.06)$ at 90° . By comparing the roughness parameters of Table 6 with those of the raw samples (see Table 3), we register a decrease of Ra by 21% at $\alpha=0^\circ$ whereas when α passes from 30° to 90° , Ra decreases by $\sim 60\%$. Concerning S_w , we observe a decrease by 6% at $\alpha=0^\circ$. For $\alpha > 30^\circ$, we notice an increase of developed surface loss with α , from 41% at 30° to 62% at 90° .

In terms of chemical etching kinetics (Figure 14), we measure a thickness loss of $249 \mu\text{m} (\pm 24)$ at $\alpha=0^\circ$, and $356 \mu\text{m} (\pm 19)$ at 30° . Then, a decrease of thickness loss is observed when increasing α , until $321 \mu\text{m} (\pm 10)$ at 90° .

The data of θ_{app} from Figure 15, obtained for all the α values, validate the positive effect of the chemical etching on surface wetting. Indeed, all the surfaces become hydrophilic ($\theta_{app} < 90^\circ$), independently of α . At $\alpha=0^\circ$, we obtain a θ_{app} of $51.7^\circ (\pm 6.3)$, corresponding to a surface energy of $85.7 \text{ mJ/m}^2 (\pm 9.4)$. We observe then an increase of θ_{app} at $\alpha=30^\circ$, to $75.5^\circ (\pm 9.5)$, with a decrease of surface energy down to $57.7 \text{ mJ/m}^2 (\pm 9.4)$. No significant differences are found between the θ_{app} at $\alpha=30^\circ$, 60° , and 90° ($p=0.84$, Kruskal-Wallis method). Compared to the raw SLM surfaces, θ_{app} decreases by about 32% at $\alpha=0^\circ$ whereas for other values of α , θ_{app} decrease by between 10% and 20%. Concerning the surface energy, we observe an increase by around 40% for all the values of α , compared to the untreated surfaces (see Figure 5).

4. Discussion

This study analyzed the effects of HF/HNO₃ chemical etching compositions and immersion times on the surface topography and wetting for SLM surfaces.

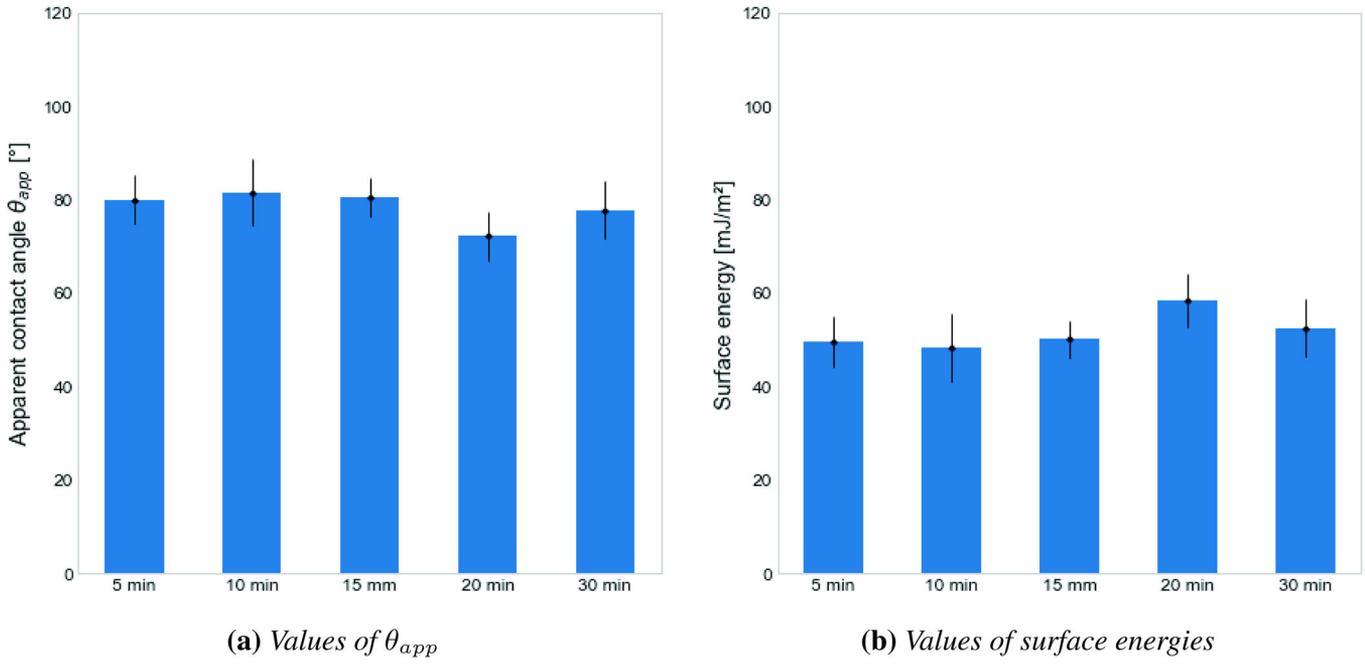


Figure 12. Values of θ_{app} and surface energies for a chemical etching of 3% HF / 10% HNO₃ on polished surface with different times of immersion.

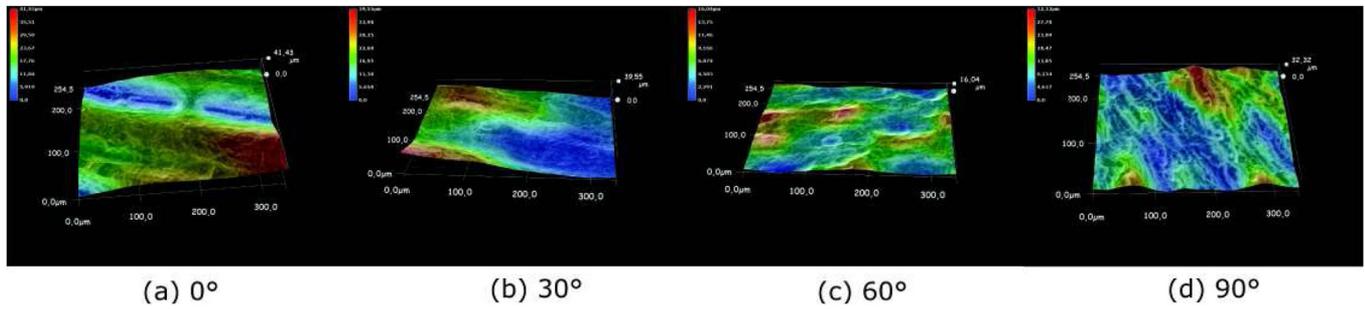


Figure 13. Surface morphologies of the surface for a chemical etching of 3% HF/10% HNO₃.

Table 6. Values of roughness parameters depending on α .

	Ra (μm)	Rq (μm)	Rz (μm)	Rsk (μm)	Rku (μm)	S_w
0°	7.88 (\pm 0.56)	9.55 (\pm 0.7)	36.24 (\pm 8.24)	-0.077(\pm 0.14)	3.08 (\pm 0.71)	1.30 (\pm 0.02)
30°	6.34 (\pm 0.34)	7.41 (\pm 0.36)	41.70 (\pm 2.23)	-0.3 (\pm 0.06)	3.11 (\pm 0.24)	1.35 (\pm 0.03)
60°	4.79 (\pm 0.21)	5.61 (\pm 0.16)	32.67 (\pm 2.7)	0.13 (\pm 0.17)	2.92 (\pm 0.28)	1.29 (\pm 0.01)
90°	6.51 (\pm 1.39)	7.76 (\pm 1.64)	40.65 (\pm 4.03)	-0.03 (\pm 0.22)	3 (\pm 0.65)	1.23 (\pm 0.06)

4.1. Remarks on surface roughness parameters

The analysis of raw SLM surface topographies pointed out significant differences depending on the printing orientation angle α . The surface roughness parameters (Ra , Rq , Rz , and S_w) increased with α (see Table 3). Especially, the increase of Rsk from negative to positive values when augmenting α proved to be a measure of increasing peak distributions due to the higher number of sintered particles trapped in the surface. These topographic variations were the results of different types of surface defects: type II defects (the melting bed texture and the staircase effect) and type I defects (the sintered particles trapped in the surface) [16]. Type I defects on the surface were assumed to cause the hydrophobic properties of the SLM surface. Indeed, excepted $\alpha = 0^\circ$, our results (Figure 5) showed θ_{app} superior to 90° for all the other orientations in agreement with the presence of type I

defects on the surface. These results agreed with different studies of the literature [41, 42].

Three different HF/HNO₃ chemical etchings had been considered, and they proved to have significant effects on the surface topography and wetting. Indeed, the surface roughness parameters showed a progressive decrease with the duration of the chemical etching for all the compositions. For the chemical etching 3% HF/10% HNO₃, we observed a continuous decrease of roughness parameters (Ra , Rq , and Rz) down to 30 min whereas the same parameters stopped decreasing at 20 min for the chemical etching 3% HF/20% HNO₃, followed by an increase of roughness (see Table 4). The chemical etching 10% HF/10% HNO₃ generated a higher roughness decrease, down to around $2 \mu\text{m}$ at 30 min. This outcome suggests that the proportion in HF controls the power of the reaction, and so the roughness reduction, whereas the proportion in HNO₃ controls

the speed of the reaction (results validated in the literature [43]). The skewness value (Rsk) was positive for the non-etched sample, due to a predominance of peaks in the surface. After the different etchings, a decrease of skewness proportional to etching duration was observed for all the different compositions. 3% HF/10% HNO_3 showed a decrease of Rsk and Rku down to 30 min where a negative value was observed, whereas negative values appeared at 20 min for 3% HF/20% HNO_3 . 10% HF/10% HNO_3 showed negative values of Rsk since the first 5 min and an increase of Rku . These results highlight that the chemical etching decreases the peak distribution density during the first phase, then it creates local porosities on the surface. The weight loss varied depending on the chemical compositions and the times of immersion (see Figure 9). For all the compositions, high weight losses were observed during the initial phase (first 5 minutes), followed by a high thickness loss.

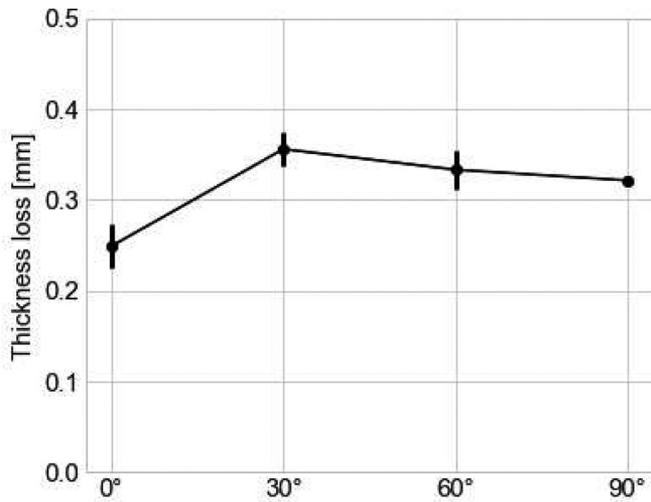
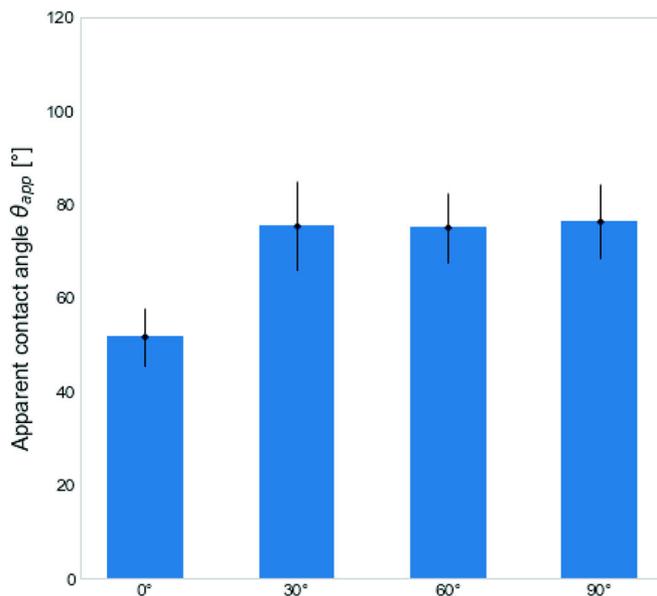
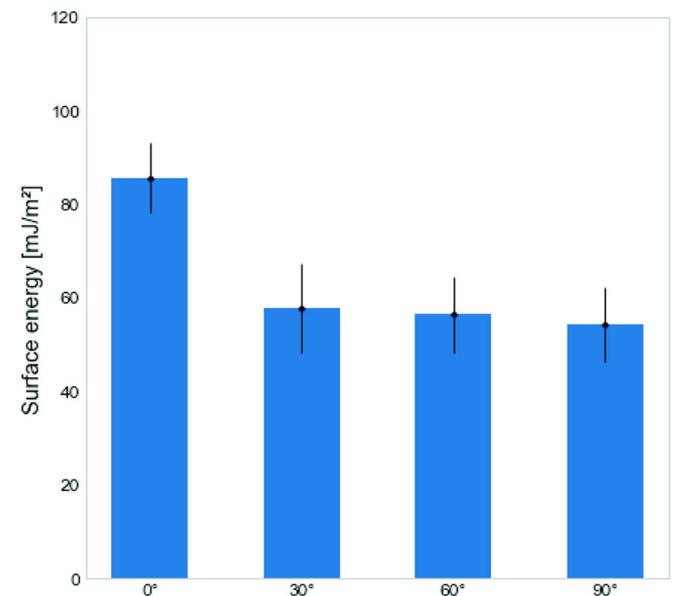


Figure 14. Thickness loss of surfaces for a chemical etching of 3% HF/10% HNO_3 .



(a) Values of θ_{app} .



(b) Values of surface energies.

Figure 15. Values of θ_{app} and surface energies for a chemical etching of 3% HF / 10% HNO_3 depending on α .

This can be explained by the high quantity of reactives at the initial state, but also by the external layer of material which is very rough (equivalent to a homogeneous layer of titanium alloy material with a lower density). 3% HF/10% HNO_3 and 3% HF/20% HNO_3 showed similar weight loss and thickness loss until $t=20$ min. After, 3% HF/10% HNO_3 showed a higher etching rate. This was reflected on the parameter S_w which decreased monotonically until 30 min (see Table 4). 10% HF/10% HNO_3 showed the highest etching rate with a corresponding weight loss of 60% and the highest decrease of S_w . In terms of wetting properties, the contact angle θ_{app} was reduced on etched samples. Indeed, for all the compositions, the initial phase ($t=5$ min) was followed by a decrease of θ_{app} and an increase of surface energy, with the disappearance of the type I defects (i.e. sintered particles) (see Figure 10). 3% HF/10% HNO_3 composition generated a decrease of θ_{app} until $t=15$ min, then the value increased and stabilized at $t=30$ min. The same behavior was observed for 3% HF/20% HNO_3 composition. However, in this last case, the minimum value of θ_{app} was reached faster than the previous one due to the higher proportion of HNO_3 . If we compare our results to the literature, the trends described beforehand match those of several published papers. Wysocki et al. [42] studied the effect of chemical etching of SLM surfaces using a solution of 2.2% HF/20% HNO_3 . The authors observed a decrease of θ_{app} from around 100° at $t=0$ min to around 45° after 6 min. In our case, the composition 3% HF/20% HNO_3 followed the same behavior, even if the reduction of θ_{app} was less pronounced. However, results relative to longer etching times have not been reported in [42]. In our study, for a time greater than 5 min, 10% HF/10% HNO_3 composition generated an increase of θ_{app} followed by a stable phase. This makes us assume that the minimum of θ_{app} was reached during the initial phase.

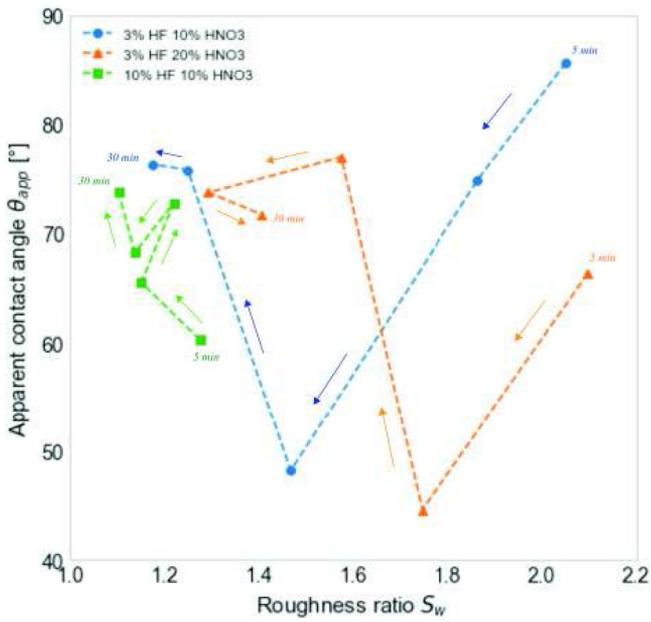


Figure 16. The relationship between apparent contact angle and roughness ratio for the SLM surfaces etched with different bath compositions (3% HF/10% HNO₃ in blue, 3% HF/20% HNO₃ in red, 10% HF/10% HNO₃ in green) at different immersion times. The arrows show the increase of the time of chemical etching.

4.2. Surface properties and wetting model

Wettability is defined as the thermodynamic interface equilibrium between the liquid, solid, and gas phases. Thus, θ_{app} variation can be provoked by the modifications of the solid phase, the two other phases being constant. These modifications can be of different nature: chemical, topographical, mechanical, etc. In particular, the roughness is proved to have a significant influence on the wettability [39, 40]. The results obtained in this study showed that the roughness is an important factor to be considered in the wetting behavior of surfaces. The initial roughness decrease was shown to be related to the elimination of type I defects (see Figure 8), causing a decrease of θ_{app} . A further decrease of θ_{app} was observed with changes in topography (see Figure 8 and Table 4), followed by a stabilization for 10% HF/10% HNO₃ or an increase for 3% HF/10% HNO₃ and 3% HF/20% HNO₃. In order to determine the evolution of wetting behavior of the surface during chemical etching, experimental data can be analyzed using wetting models taken from literature. The most classical model evaluating the apparent contact angle as function of the roughness is the Wenzel model [44]. According to this model, the roughness is considered through the introduction of the “roughness ratio” as the ratio between the actual surface and the projected one. Thus, the apparent contact angle θ_{app} is calculated as the product between the the roughness ratio S_w and the intrinsic contact angle θ_γ :

$$\cos \theta_{app} = S_w \cdot \cos \theta_\gamma \quad (4)$$

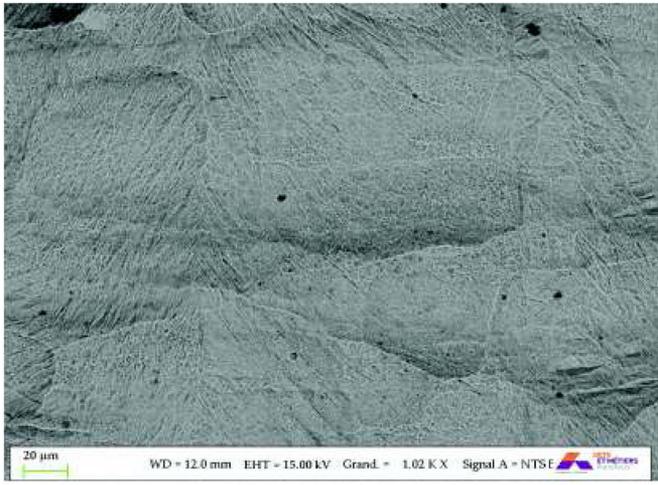
The Wenzel model, however, applies only when a perfect contact is ensured between an homogeneous solid surface and the liquid. Since for a given solid/liquid material combination the intrinsic contact angle θ_γ does not depend on

S_w , eq. 4 states that when $\theta_\gamma < 90^\circ$, the higher the roughness (thus the higher S_w), the lower θ_{app} . On the contrary, if $\theta_\gamma > 90^\circ$, the higher the roughness, the higher θ_{app} .

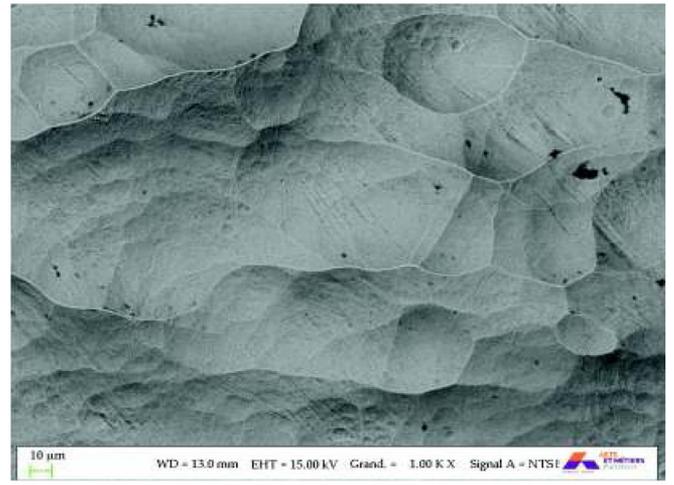
Concerning the wetting nature of flat TA6V surfaces, the experimental measurements suggest that the surface is hydrophilic, see [45, 46]. However, beyond the surface roughness effect, the SLM process as well as the chemical etching could let vary material properties and chemical composition of the surface, respectively. Indeed, concerning SLM process, an analysis of the material behavior has shown the presence of a gradient of hardness (Figure 6) from the external border of the specimen to the internal core. This material gradient was followed by a gradient of wetting properties at different depths within the specimen (see Figure 7), more precisely, with a decrease of θ_{app} from the external border to the internal core. This phenomenon could be the result of a faster cooling time of the material on the borders compared to the one of the core. Several works of literature had already shown an increase of hardness on materials submitted to slower cooling process [47], via a heat treatment, followed by a decrease of θ_{app} compared to the material without heat treatment [41]. Finally, it can be assumed that the SLM process is not a factor that could change the wetting properties of the surface toward hydrophobicity. On the other hand, concerning chemical etching effect in terms of chemical composition of the solid surface, titanium surfaces treated by acid hydrofluoric HF have shown changes in their wetting properties. Zahran et al. studied the wetting properties of titanium surfaces treated with HF. They explained that if the chemical changes affected the decrease of θ_{app} during the first 3 min, the most influencing parameters for longer immersion times were the roughness parameters via the form and the distribution of the peaks. Taking into account this literature results, we can assume that flat TA6V surfaces obtained by SLM and treated chemically exhibit hydrophilic properties.

Figure 16 presents the evolution of θ_{app} as a function of S_w and at the five different immersion times (5, 10, 15, 20 and 30 min). The straight lines and arrows indicate the direction of increasing immersion time. For each one of the applied chemical compositions, a global decrease is observed for low values of S_w which correspond to high values of immersions time. These parts follow the trend provided by the Wenzel’s model. Moreover, this result is supported by the fact that at those values of immersion time the variation of chemical composition of the surface is negligible. On the other hand, in the case of 3% HF/10% HNO₃ and 3% HF/20% HNO₃, an increase of θ_{app} is observed for high values of S_w (i.e. beyond 1.41 and 1.75, respectively). This values of S_w correspond to brief immersion times wherein the chemical composition variation of the surface could be considerable. This suggests that, at those stages, the θ_γ could vary due to chemical composition variation. Otherwise, these results mean that the Wenzel’s model is no longer applicable and the hypothesis of full contact is not valid at high S_w .

If the Wenzel’s model no longer applies, the influence of the roughness can be described by a more general model:



(a) 3% HF / 10% HNO₃



(b) 3% HF / 20% HNO₃



(c) 10% HF / 10% HNO₃

Figure 17. SEM image of the surface condition in BSD for a chemical etching of 15 minutes.

the Cassie-Baxter's model [48]. In the Cassie-Baxter's model, the drop is assumed to not fully fill the surface and that small pockets of air stay between the drop and the surface. Thus, we obtain a mix of solid-liquid-gas contact surface satisfying the thermodynamic equilibrium. Thus, Cassie and Baxter defined a model relating θ_γ to θ_{app} via the fractional area of solid material f_s with respect to the whole contact area [48], expressed as:

$$\cos \theta_{app} = f_s \cdot (1 + \cos \theta_\gamma) - 1 \quad (5)$$

It can be assumed that an increase of the roughness results in the increase of air pockets on the surface. Thus, this decreases the fractional area f_s . Therefore, the higher the roughness (thus the higher S_w), the higher θ_{app} . This behavior, predicted by Cassie-Baxter's model, is in agreement with the results of 3% HF/10% HNO₃ and 3% HF/20% HNO₃ at high values of S_w (i.e. equal or higher to 1.41 and 1.75 respectively). This means that, very probably, at high S_w some porosities appear at the solid/liquid contact surface.

Thanks to results of Figure 16 and taking into account literature results, we can assume that different wetting regimes

take place on SLM surfaces during the chemical etching. If we consider the chemical composition 3% HF/10% HNO₃, we observe a first phase at $S_w > 1.41$ where the roughness is very probably (since no porosity observation has been conducted) high enough to trap air pockets at the solid/liquid interface. Thus, the Cassie-Baxter's model applies with a decrease of the apparent contact angle as the roughness decrease. For $S_w < 1.41$, the air pockets leave the surface as the roughness decreases. Thus, the wetting regime follows the Wenzel's model with an increase of the contact angle as the roughness decrease. The same behavior can be observed on 3% HF/20% HNO₃ chemical composition, with no significant differences observed for $S_w < 1.5$. The chemical composition 10% HF/10% HNO₃ globally follows the Wenzel's model, with the presence of the speak due to the measurement variability. Some studies propose further models for the case of wetting regimes switching from the Wenzel's model to the Cassie-Baxter's one. We can mention the model of Nagayama et al. [49] that proposed a partial wetting model for nano/microstructured surfaces. However, for the sake of brevity, no further models will be analyzed in the present work.

4.3. Remarks on biocompatibility

To conclude, in order to remove type I defects (sintered particles), all analyzed chemical compositions are sufficient with an immersion time lower or equal to 5 min. In terms of wettability, 3% HF/10% HNO₃ composition generated the highest surface energy after an immersion time of 15 min (see Figure 10). 10% HF/10% HNO₃ gave a lower value of the maximum of surface energy but we could assume that the maximum of surface energy was reached before the 5 min. However, its surface energies gave better values than the two other compositions for longer times. The surface roughness of 10% HF/10% HNO₃ composition was the fastest one to reduce roughness and gave the lowest roughness values ($R_a = 2 \mu\text{m}$ at $t = 30 \text{ min}$). On the contrary, when using 3% HF/20% HNO₃, the minimum value of R_a was $7.9 \mu\text{m}$ and increased again after some time. 3% HF/10% HNO₃ composition generated a minimum value of R_a lower than $7 \mu\text{m}$, with a homogeneous etching. This composition improved surface quality and wettability of all surfaces at different α (see Figures 13 and 15, and Table 6). Only 10% HF/10% HNO₃ composition for an immersion time of 30 min filled the objectives of the work in terms of surface roughness. However, this set of parameters presents challenging issues: (i) The weight loss reached to 60% with more than 20% of weight lost during the first 5 minutes, corresponding to a thickness loss of $400 \mu\text{m}$. This effect has to be taken into account in the design of the component in terms of tolerance dimensions. (ii) The reaction was really powerful and caused the creation of a lot of NO₂ gas in the solution. This could be a problem for complex geometries where the gas could stuck on the internal surface and prevent a homogeneous etching [37]. In the same way, the reaction can be heterogeneous and create a lot of porosity on the surface. (iii) We observed the creation of a layer of oxide after a certain time. It can be a hindrance for biomedical applications. (iv) The reaction created cracks on a certain number of samples. This can be the combination of the powerful reaction with the residual stress, and thus can be prevented by heat treatment. Besides, if we look at 3% HF/10% HNO₃ and 3% HF/20% HNO₃ surfaces, we noticed that at the sub-micron scale, the surface were relatively flat (see Figure 17) and that most of the roughness was present at a sub-meso scale. So, if we compare to other studies in the biomedical field where surface energy and wetting are drastically altered on nanometer (less than 100 nm) and sub-micron (greater than 100 nm) scales [30], those surfaces could be considered as viable.

5. Conclusion

In the present work, the effect of the chemical etching to improve the topography and the wettability of SLM-based surfaces has been analyzed. First, raw SLM surfaces of samples printed at different inclinations α (with respect to the printing plateau) have been observed. Moreover, the gradient of material properties of specimens (due to SLM process) has also been evaluated. Then three different surface treatment protocols have been applied on both raw and

polished samples. The results showed that raw SLM surfaces orientated at $\alpha > 0^\circ$ exhibited hydrophobic properties, due to sintered particles (type I defects) trapped into the surface. The analysis of surfaces treated with three chemical etching compositions and five different immersion times have shown an elimination of type I defects followed by an increase of wetting properties. The three etching compositions reached a minimum apparent contact angle θ_{app} for different immersion times. The analysis of the surface wetting showed that the key factor to improve this property is the surface roughness. The comparison of literature wetting models provided some explanations on the influence of the roughness as function of chemical etching composition and times. Furthermore, there are also second-order factors. Indeed, surface wetting is also influenced by the variation of material properties. The biomedical application of integrated implants has been selected as target to attain specific ranges of surface roughness and wettability (through the proposed surface treatments). Results show that the main requirements of wetting were met but not the conditions of average roughness. Therefore it can be assessed that chemical etching is a suitable surface treatment to improve biocompatibility of SLM-based implants. However further studies could let analyze similar chemical compositions in order to reach simultaneously both requirements on surface wetting and average roughness.

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