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Evolution of residual stresses during short time nitriding of 33CrMoV12-9 steel grade.

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Abstract

The evolution of residual stresses during short time gas nitriding of 33CrMoV12-9 steel grade is studied. It aims understanding the influence of nitriding parameters (temperature T and nitriding potential Kn) on the generation and evolution of residual stresses in the very first stage of nitriding. The samples are gas nitrided using a thermobalance during 2h30 and 5 hours for various temperatures and nitriding potentials. Residual stress analyses are carried out by laboratory X-ray diffraction.

Introduction

Nitriding is a thermochemical surface treatment involving the diffusion of nitrogen and generating a hardening effect combined to compressive residual stresses [1]. This process is used to improve fatigue life, tribological and anti-corrosion properties of iron-based alloys [2]. At nitriding temperature (400 to 600°C), the generation of residual stresses occurs simultaneously with hardening. It finds its origins in volume changes accompanying phase transformations during nitrogen diffusion. The volume change is a consequence of lattice distortions, molar volumes and thermal expansion coefficients between precipitates and the ferritic matrix [3-4].

The influence of nitriding parameters (temperature and nitriding potential) on the evolution of residual stresses has been studied by Barralis [4] for long time gas nitriding of 33CrMoV12-9 steel grade and by Jegou [6] on a ternary Fe-3%Cr-0.345%C alloy. The evolution of residual stresses during the nitriding treatment is the consequence of the microstructural evolutions through the case due to the co-diffusion of nitrogen and carbon. The resulting precipitation kinetic modifies the volume fraction of precipitates as well as the induced volume variation [7].

The present work aims studying the generation of residual stresses in the very first stage of nitriding of the 33CrMoV12-9 steel grade. Various nitriding parameters (T and Kn) are considered.

Experimental Procedure

Material, sample preparation and nitriding process

The composition of the 33CrMoV12-9 steel grade, quenched and annealed at 620 °C, used during this study is given in Table 1. Before gaz nitriding, rectangular specimens (17x13x5 mm³) are machined on all sides and cleaned by ultrasounds within ethanol.

Nitriding experiments were performed in a Setsys Evolution thermobalance with a vertical alumina-tube furnace controlled within ± 3 °C in the isothermal zone. The nitriding atmosphere was composed of a gas mixture of NH₃-N₂-H₂. The flow rate of each gas was adjusted separately using mass-flow controllers. The nitriding potential Kn is given by the composition of gas mixtures as follows (eq.1):

$$Kn = P_{NH_3}/(P_{H_2})^{3/2} \quad \text{eq 1}$$

The nitriding parameters of each sample are given in Table 2. Figure 1 shows the position of each treatment on the Lehrer diagram [8]. In case of pure iron, it gives the phase in equilibrium with the nitriding gas atmosphere depending on the temperature T and potential Kn . The dashed line on the Lehrer diagram corresponds to the interface between α -Fe and γ' -Fe₄N as experimentally obtained for the 33CrMoV12-9 grade steel (see Fig.1) [8]. After nitriding, the specimens were cooled at 20 °C.min⁻¹ to room temperature in a flow gas of nitrogen.

The nitrided samples were cut in two pieces, grinded and polished up to a 0.05 μm colloidal silica suspension. The nitrided cases were observed by scanning electron microscopy using a backscattered electrons mode (JEOL 7001F SEM-FEG).

Stress analysis

Residual stress analyses were carried out by laboratory X-ray diffraction using a Siemens D500 diffractometer with Cr-K _{α} radiation ($\lambda = 0.228975$ nm) on the {211} diffracting plane of α -Fe. The gauge volume was defined by a 3 mm diameter tape mask. The $\sin^2\Psi$ method was used to determine the mean residual stresses in order to determine $\sigma^a_{XRD} = (\sigma^a_{xx} - \sigma^a_{zz})$ in the ferritic matrix within the depth of the case. Radiocrystallographic stiffness are calculated with Kröner-Elshelby model and used to determine the stress ($S_1^{\{211\}} = -1.25 \cdot 10^{-6}$ MPa⁻¹, $\frac{1}{2} S_2^{\{211\}} = 5.85 \cdot 10^{-6}$ MPa⁻¹). Depth profiling was performed using electro-chemical surface layer removal (Struers Lectropol 5).

Table 1: Composition of the 33CrMoV12-9 steel grade.

wt.%	C	Si	Mn	Cr	Mo	Ni	V	Fe
33CrMoV12-9	0.34	0.23	0.45	3.10	0.91	0.22	0.21	Balance

Table 2: Nitriding time, temperature and potential for gas nitriding of 33CrMoV12-9 steel grade.

Time [h]	Temperature [°C]	Kn [atm ^{-1/2}]
5h	550	0.30
5h	550	0.79
5h	550	3.65
5h	550	13.77
5h	520	3.65
5h	480	3.65
2h30	550	3.65

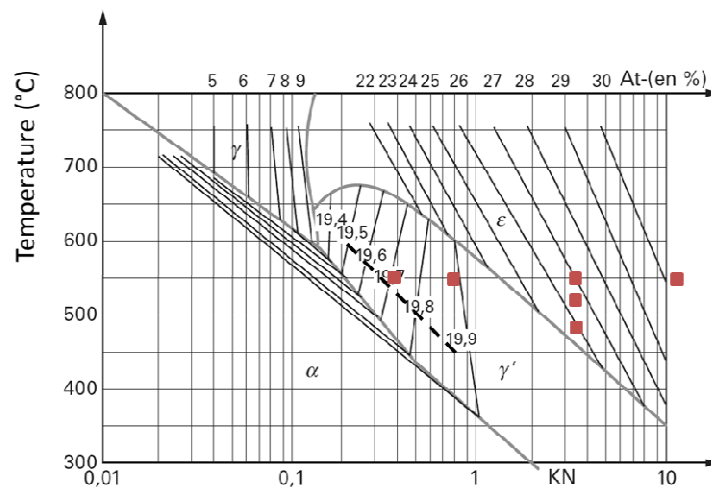


Figure 1: Lehrer diagram [8] and nitriding parameters of Table 2 as red points. The dashed line corresponds to the interface between α -Fe and γ' -Fe₄N as experimentally characterized.

Results and Discussion

Influence of the nitriding potential

SEM micrographies (Figure 2) show that the thickness of the compound layer and precipitation of cementite at grain boundaries nearly parallel to the nitriding surface increase with the nitriding potential Kn . According to the chemical contrast, iron nitrides are also observed at grain boundaries at the interface between the compound and diffusion layer.

The in-depth stress profiles are given on Figure 3. The maximal compressive stress is of the same magnitude for all nitriding potentials after 5 h of nitriding at 550 °C. The surface zone under compressive stresses increases with the nitriding potential, being almost similar for $Kn = 0.79$ and 3.65. The evolution of the compressive residual stress of the near surface is remarkable even after only 5 h of nitriding. Between 0 and 50 μm in depth, residual stresses (taking the absolute value) change with the nitriding potential. For a given phase region of the Lehrer diagram, but decrease when changing from the γ' -Fe₄N to the ϵ -Fe₂₋₃N region. These trends are similar to those found by Barralis for long time nitriding of the 33CrMoV12-9 grade. Table 3 show the Specific volume of the phase present in nitriding steel.

According to Dulcy *et al.*, a low nitriding potential ($0.4 < Kn < 2$) induces a decarburization of the surface by the atmosphere whereas a high nitriding potential ($Kn > 2$) induces a diffusion of carbon toward the core of the sample [8]. The effect of carbon diffusion on the generation and evolution of stress during nitriding were studied in previous works [6, 7]. The precipitation of cementite at grain boundaries involves positive volume change (see Table 3) generating compressive residual stresses. During the transformation of cementite into γ' -Fe₄N nitride, the residual compressive stresses decreases (corresponding to the first point of the residual stress profile $Kn = 3.65$). The transformation of cementite into γ' -Fe₄N nitride induces positive volume change that generate compressive residual stresses (corresponding to the first point of the residual stress $Kn = 0.79$). The difference of the compressive stress near the surface between samples can thus be explain by different volume fraction of cementite and different stage of transformation from cementite into iron nitride. The positive volume change is higher for the transformation of cementite into γ' -Fe₄N. The increase of residual stresses just after the maximum of stress (depth > 100 μm) when increasing the nitriding potential suggests different precipitation kinetics due to the diffusion of nitrogen but also carbon. The diffusion of carbon toward the nitrogen diffusion front is more important for the sample nitrated at $Kn = 13.77$. It results in the precipitation of carbides M₇C₃, M₂₃C₆ or cementite depending on the content of nitrogen [7]. Such precipitation generates positive volume change and results in the generation of compressive stresses. According to Dulcy *et al.*, the nitrogen concentration at the interface γ'/α varies from 1.4 % to 1.86 % at 520 °C for a Kn going from 0.4 to 1.62 in the 33CrMoV12-9 steel grade [8]. This change in concentration, and there for the corresponding increase of the nitrogen flux, can explain the increase of the size of the surface zone under compressive stress between $Kn = 0.3$ and $Kn = 0.79$.

Table 3 : Specific volume of phases encountered within a nitrated case (M = Cr, V, Mo...) [8].

Phase	Specific volume (cm ³ .kg ⁻¹)
α -Fe	127.1
M ₃ C	130.9
γ' -Fe ₄ N	138.1
ϵ -Fe ₂₋₃ N	138.7
Carbides (M ₇ C ₃ M ₂₃ C ₆)	145.5
Nitrides (CrN)	161.3

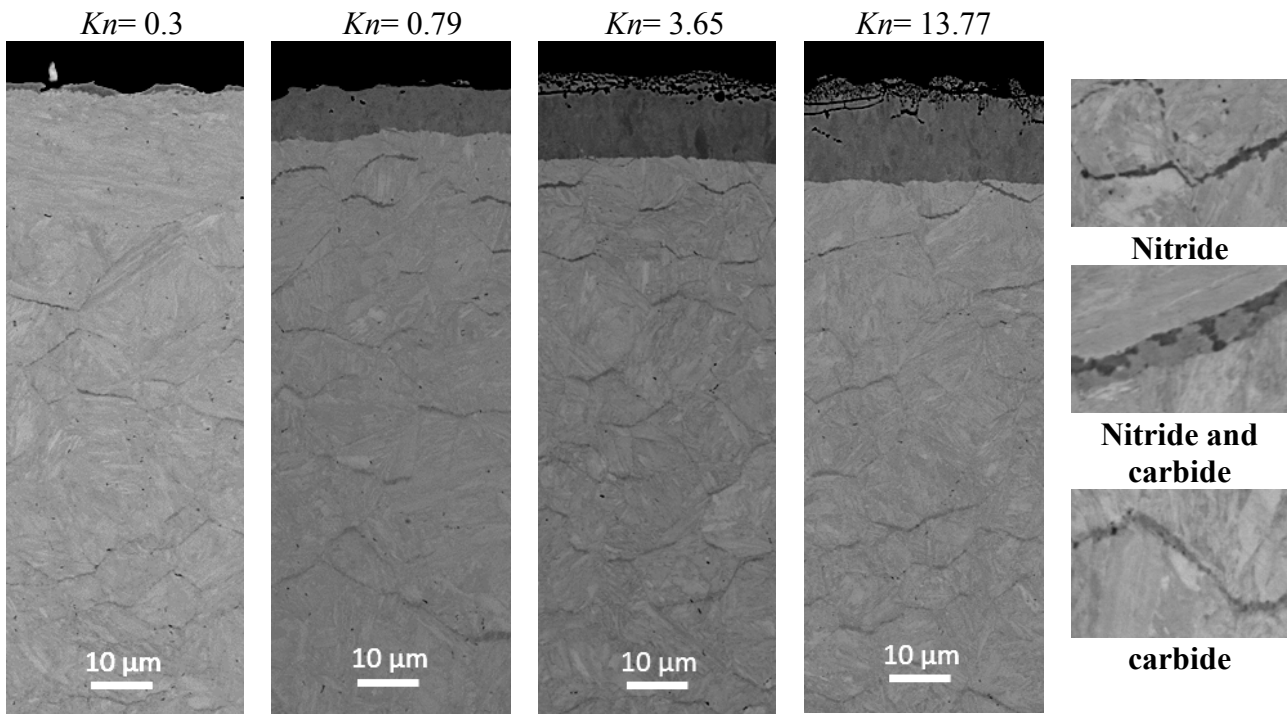


Figure 2 : SEM micrographies using backscattered electrons of the 33CrMoV12-9 steel grade nitrided at 550 °C during 5 h.

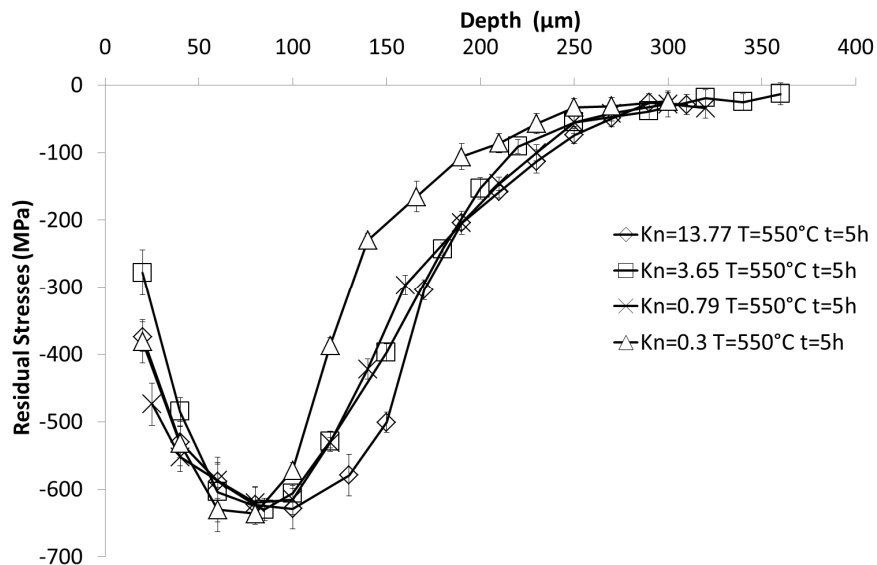


Figure 3: Influence of the nitriding potential on the mean residual stress of the ferritic matrix. Depth profiling of the 33CrMoV12-9 steel grade nitrided at 550 °C during 5 h.

Influence of the nitriding time

Figure 5 gives the evolution of residual stresses in the ferritic matrix with the time of nitriding at 550 °C and for $Kn = 3.65$. In this case the maximal compressive stress decreases with the nitriding time. As could be expected, the depth affected by compressive residual stress increases with the treatment time. As previously, the compressive residual stress close to the surface decrease when the time increases due to the transformations of cementite into γ' -Fe₄N. This evolution in the stress profile with the treatment time can be attributed to the co-diffusion of nitrogen and carbon [7]. The

diminution of the residual stress with the time can be due to an evolution of the precipitation during the nitriding.

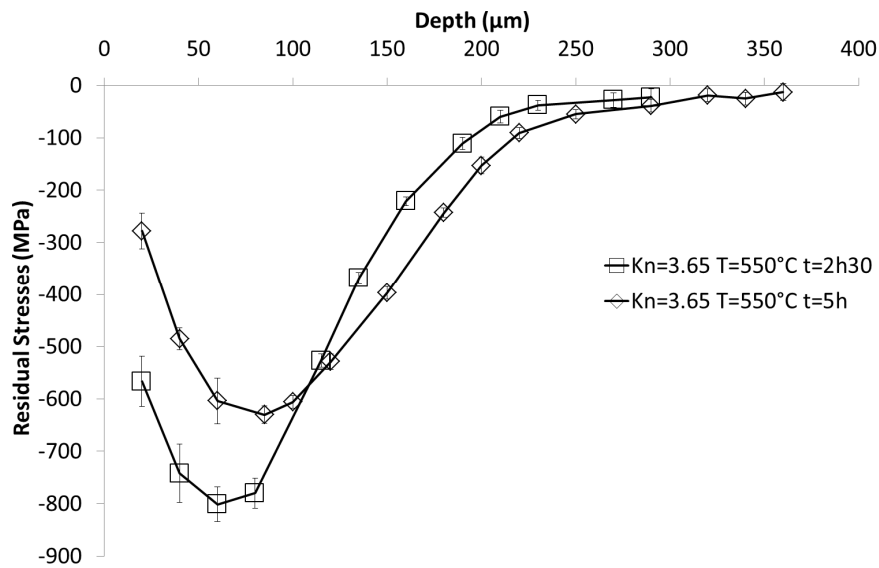


Figure 5: Influence of the nitriding duration on the mean residual stress of the ferritic matrix. Depth profiling of the 33CrMoV12-9 steel grade nitrided at 550 °C for $Kn = 3.65$.

Influence of the nitriding temperature

Figure 4 shows the in-depth evolution of residual stresses of the ferritic matrix with the nitriding temperature. Residual stresses (taking the absolute value) close to the surface highly decrease with the temperature, going from 1600 MPa at 480 °C to 250 MPa at 550 °C. The maximal compressive stress decreases with the temperature. The surface zone under compressive stresses increases with temperature. This co-diffusion induces phase transformations, accompanied by positive volume change and the generation of compressive residual stresses. The evolution of the maximal compressive stress is due to a thermal relaxation and a change in the kinetics of phase transformation. As shown for long time nitriding, such evolution is characteristic of a change in the precipitation kinetics induced by the co-diffusion of nitrogen and carbon [7]. These results are similar to those found by Barralis for long time nitriding of the 33CrMoV12-9 steel grade.

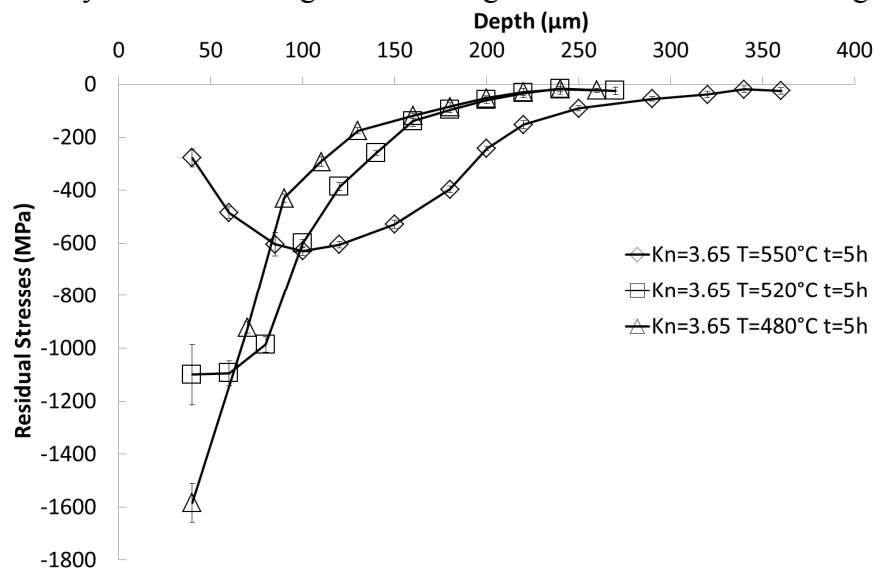


Figure 4: Influence of the nitriding temperature on the mean residual stress of the ferritic matrix. Depth profiling of the 33CrMoV12-9 steel grade nitrided during 5 h for $Kn = 3.65$.

Conclusion

The influence of nitriding parameters (time, temperature, nitriding potential) on the residual stress generation and evolution was investigated for a nitrided 33CrMoV12-9 steel grade.

The nitriding potential modify the carbon distribution of the nitrided layer that increase the volume fraction of cementite precipitation. Increasing the nitriding potential beyond $Kn = 3.65$ permits to increase the surface zone under compressive stresses. The nitriding potential have no influence on the value of the maximum compressive stress.

The nitriding temperature activates the diffusion of nitrogen and carbon and accelerates the kinetics of phase transformations. Modifying the nitriding temperature permit to control the value of the maximum compressive stress and the size of the zone under compressive stress.

The nitriding time modify the value of residual stress close the surface, the value of the maximum compressive stress and the size of the under compressive stress.

The results found for short time nitriding are different to those found by Barralis for long time nitriding of the 33CrMoV12-9 steel grade [5]. However to explain the influence of nitriding parameters, more investigations are required in order to quantify the diffusion of carbon as well as the volume fraction of iron nitrides and cementite.

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